Characterization and Implications of Solids Associated with Hydraulic Fracturing Flowback and Produced Water from the Duvernay Formation, Alberta, Canada

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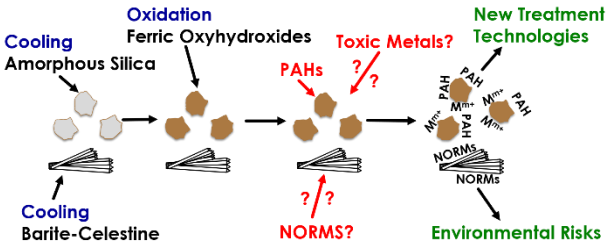
Abstract:

Public concern is heightened around flowback and produced water (FPW) generated by the hydraulic fracturing process. FPW is a complex mix of organic and inorganic solutes derived from both the injected hydraulic fracturing fluid and interactions with the subsurface lithology. Few studies to date have systematically investigated the composition of FPW or its individual components. Here, we provide the first systematic characterization of the composition of the solids associated with FPW by analyzing samples from three wells drilled into the Duvernay Formation in Alberta, Canada. The FPW initially returned to the surface with high total dissolved solids (greater than 170,000 mg/L) and enriched with Fe(II), silica, sulfate, barium, and strontium. The solids form two distinct phases once the FPW reached the surface: (1) silica-enriched Fe(III) oxyhydroxides, and (2) a barite-celestine solid solution. We hypothesize that the precipitation of the amorphous silica-enriched Fe(III) oxyhydroxide is a two-step process, where first the silica precipitates as a function of the cooling of the FPW from elevated subsurface temperatures to ambient surface temperatures. Next, the silica acts as a template for the precipitation of Fe(III) oxyhydroxide as the diffusion of oxygen into the subsurface causes oxidation of aqueous Fe(II). The barite-celestine solid solution precipitates solely as a function of cooling. Elevated dissolved Fe concentrations in FPW and modeled saturation indices from five North American shale plays (Marcellus, Fayetteville, Barnett, Bakken, and Denver-Julesburg) indicates that solids similar to those found in Duvernay FPW, specifically Fe(III) oxyhydroxides, barite and quartz, are likely to occur. With the solids known to carry a significant portion of FPW's toxicity and organic contaminant load[1], the development of new treatment technologies, such as the oxidation of the Fe(II) in FPW, may increase FPW reuse and reduce the environmental risk posed by FPW.

40 

41 **Key Words**

42 Hydraulic fracturing, Flowback and produced waters, Iron-oxides, Treatment

43

1 Introduction

Hydraulic fracturing for oil and gas extraction has generated a range of environmental concerns including air pollution[2], induced seismicity[3-5], fresh water allocations[6-8], drinking water contamination[9-10], and human health risks[11-13]. While there has been substantial public concern on the potential for fluids related to hydraulic fracturing migrating to the shallow subsurface and contaminating groundwater, the most likely routes for groundwater contamination are the loss of wellbore integrity or surface spills during the handling and transport of hydraulic fracturing fluid (HFF) or flowback and produced waters (FPW)[9,14-18].

Much of our current understanding of the composition of FPW comes from just a few locations, namely the Marcellus Formation (Northeast USA)[19-21], Julesburg basin (Colorado, USA)[22-23], Duvernay Formation (Alberta, Canada)[1], and Horn River Formation (British Columbia, Canada)[24]. Generally, FPW has been found to contain high total dissolved solids (TDS), ranging from 10,000 to over 300,000 mg/L depending on the formation, fracturing process, and time. The TDS is largely composed of Na and Cl, with elevated concentrations of Ca, Fe, Mg, and Sr [20-24]. Trace elements, such as Se, As, Ba[19], and naturally occurring radioactive materials (NORMs)[25-26], have also been documented in FPW.

The organic chemical composition of FPW is also complex. FPW is composed not only of the hydrocarbons naturally present in the target geologic formation, but also HFF chemicals and the compounds that form as a product of secondary reactions when HFF is exposed to elevated subsurface temperatures and pressures[1,27]. As FPW can contain numerous combinations of the thousands of chemicals known to have been used in hydraulic fracturing fluids alone[13], targeted analysis can be difficult. To date, a variety of compounds have been either identified or

quantified in FPW, including volatile organic compounds (e.g., benzene, toluene, ethylbenzene, and xylene)[27-29], polymers (e.g., polyethylene glycols and polypropylene glycols)[30-32], biocides (e.g., glutaraldehyde and alkyl dimethyl benzyl ammonium chloride: Ferrer and Thurman, 2015), surfactants (e.g., dimethylamines, cocamidopropyl and its derivatives)[33], and reaction byproducts (e.g., halogenated methanes and acetones, chloromethyls, delayed acids, and organophosphates)[1,27].

FPW compositional studies have focused the analysis primarily on the fraction that passes through a membrane (0.2 or 0.45 μm), i.e., the dissolved phase. With the exception of a few filtration studies[34-37], such as Phan et al.[38] who examined the composition of colloidal fractions that passed through 0.45 μm pore filters, the associated solid and colloidal fractions that do not pass through the filters have been largely ignored. This is problematic because in previously studied contaminated aquifers, colloidal fractions are a major vector for the transport for contaminants, such as uranium[39-40], plutonium[41-42], and polycyclic aromatic hydrocarbons (PAHs)[43-44]. Indeed, recent work has demonstrated strong links between the FPW solids, PAHs and environmental toxicity[1,45-46]. Specifically, He et al.[1] showed that zebrafish (*Danio rerio*) embryos exposed to unfiltered FPW samples had approximately double the ethoxyresorufin-O-deethylase (EROD) activity compared to those exposed to the corresponding filtered samples. Additionally, the 96 h LC_{50} concentration for samples containing solids was approximately 50% lower than the corresponding samples without solids. Further, He et al.[45] found a significant increase in EROD activity and thiobarbituric acid reactive substances in gill, kidney and liver tissues of juvenile rainbow trout (*Onorhynchus mykiss*) when FPW solids were present. The exact mechanisms by which the solids cause the increased toxicity has yet to be determined,

including whether the toxicity is derived primarily from the solids themselves or the associated organic compounds.

In this study, we have characterized the chemical composition and mineralogy of FPW associated solids from the Duvernay Formation. The Duvernay Formation is a Devonian-aged formation composed of silt-sized quartz, mica, K-feldspar, and plagioclase; the clay-size fraction is dominantly quartz and calcite with subordinate clay illite/smectite and less abundant kaolinite in descending order of their relative abundance[47,48] (for more information, a full description of the geologic context is provided in the Supplemental Information). Using the FPW data we model the saturation indices for the minerals identified in the characterization of the solids and then discuss the mechanisms underpinning their formation with a comparison to the potential for secondary mineral formation in five other North American shale plays. Analysis of bulk chemistry of aqueous and solid phases capture on 0.2 µm membranes were conducted, and particle morphology and spot elemental and mineralogy of the solids were determined using spectroscopy. 0.2 µm membranes were selected as it has been suggested that 0.2 or 0.45 µm are the most commonly used pore sizes used microfiltration for FPW reuse[49]. The strong correlations observed between toxicity, oxidative stress, endocrine disruption, and the FPW solids[1,45] indicates the importance of characterizing these solids. Insights into their composition and formation will improve the handling, treatment, and disposal technologies to mitigate risks.

2 Methods

2.1 Flowback Samples

Samples of FPW from three hydraulically fractured wells located near Fox Creek, Alberta, Canada in the Devonian-aged Duvernay Formation were acquired from Encana Corporation. All samples were collected after the oil and gas separator into polypropylene buckets. Sample 1 (S1) was collected 7 days after the initiation of the flowback and collected in August 2014, which was previously examined for aqueous toxicity and characterized in He et al.[1,45]. Sample 2 (S2) collected 10 days after the initiation of the flowback and collected in November 2015[47]. Sample 3 (S3) from 24 hrs after the initiation of the flowback and collected in October 2016. Upon arrival the pH of each fluid was measured using a Mettler Toledo Easy Five dual pH meter, which was calibrated using a 3 point calibration method using buffers pH 4.01, 7.00, 10.01. The buffers were not matched to the ionic strength of the FPW and there for should be viewed as an estimate. A portion of S3 was further treated by vigorously bubbling it with air for 4 hours, subsequently referred to as S3_Ox.

2.2 X-ray Diffraction Spectroscopy

Two samples of Duvernay shale collected as drill cuttings from 700 m apart alone the horizontal bore of the hydraulic fractured well from which S3 were acquired were analyzed for their bulk mineral composition (Figure SI.1). Samples were soaked for 1 hrs in with OptimaTM grade n-Hexane at a ratio of 10mL solvent to 0.5g sample to remove residual drilling fluid organic compounds. The cuttings were recovered by centrifugation and then the remaining solvent was allowed to evaporate before analysis. Bulk minerology was determined using X-ray diffraction (XRD) using a Rigaku Ultima IV. A Co-60 electron tube radioactive source and a scan rate of 2q/min with a step size of 0.02q were used, spectra was then fit using JADE 9 software and ICDD (International Centre for Diffraction Data) and ICSD (Inorganic Crystal Structure Database).

2.3 Digestions

Samples of FPW solids for digestion were prepared using a method similar to He et al.[1]. Each of the three FPW samples were thoroughly shaken before being passed through a nylon membrane filter with a nominal pore size of 0.2 µm. The solids retained by the filter were then rinsed with 18MΩ ultrapure water while still on the filter, then covered and air dried for 48 h at room temperature.

The bulk elemental composition of the FPW solids was determined by alkaline fusion digestion preformed in triplicate using a method similar to that described by GBC Scientific Equipment[51]. Alkaline fusion digestions were chosen over hydrofluoric acid digests due to the potential for the loss of Si from the evaporation of SiF_4 and SiF_6 and the precipitation of Al, Fe, Sr fluoride minerals in hydrofluoric acid digestions. Additionally, von Gunten et al.[52] found alkaline fusion and hydrofluoric digestions to yield comparable results for total digestions. For each, 1.5 g of NaOH was melted in a nickel crucible over a Bunsen burner, and then 0.1 g of sample was added to the melt with 0.5 g of Na_2O_2 and an additional 0.5 g of NaOH. The crucibles were then covered with nickel covers and heated for an additional 5 minutes. After cooling, the crucibles were rinsed with 6 M HCl before being placed into a polypropylene beaker along with the lid. Aliquots of 6 M HCl were added until all the remaining particles dissolved (approximately 20-30 mL), the final the volume was adjusted to 100 ml with 18 MΩ ultrapure water. Solutions were diluted 10-fold with a mixture containing 2% HNO_3 and 0.5% HCl prior to analysis by ICP-MS/MS (see Section 2.3.1). Metal recoveries were verified using STSD-3 reference material (CANMET Mining and Mineral Sciences Laboratories), which was digested in parallel to the samples.

2.4 Aqueous Analysis

The aqueous phase of samples S1 and S2 were previously analyzed and reported in He et al.[1] and Blewett et al.[50], respectively. For S3 and all aqueous phase analyses, the samples were vigorously shaken before an aliquot was removed and filtered through a 0.2 µm nylon membrane filter into polypropylene ICP tubes. The filtrate was then used for all subsequent analyses.

2.4.1 ICP-MS Analysis

The aqueous phase of S3 and digestion solutions from S1-3 and S3_Ox were analyzed using an Agilent 8800 ICP-MS/MS. S3 was diluted with 18 MΩ ultrapure water with a dilution factor (DF) of 850 for Na analysis and 85 for all other elements After dilution, S3 was acidified with 12 µL of trace metal grade nitric acid per 10 mL of sample and all blanks were prepared in a similar manner with the sample fraction being replaced with 18 MΩ ultrapure water. During analysis, the RF power was set to 1550 W and the reflected power was 18 W. Samples were aspirated with a micromist nebulizer and nickel/copper cones were used. MS/MS mode was used for sample analysis to acquire greater mass resolution, and the gas collision/reaction cell was utilized with either He, O_2 or H_2 gas depending on the element. To account for instrumental drift, a standard solution of 0.5 ppm indium (In) was added to each sample utilizing an inline addition system.

2.4.2 Ion Chromatography

Ion chromatography was used to determine the chloride (Cl^-) and sulfate (SO_4^{-2}) concentrations of S3. Filtered samples were diluted 1:2000 with 18 MΩ ultrapure water and then analyzed using a Dionex Ion chromatography DX 600 with a 4mm analytical column (AS9-HC), guard column (AG9-HC), and a 4 mm ASRS Ultra suppressor.

2.4.3 Total Organic Carbon and Total Nitrogen Analysis

Total organic carbon (TOC) and total nitrogen (TN) were determined using a Shimadzu model TOC-V-CHS/CSN TOC analyzer. TOC was measured as nonpurgeable organic carbon. TN was measured as NO_2 after the sample was combusted.

2.4.4 Alkalinity Measurements

The alkalinity was determined for S1- S3 upon arrival from the field through an end-point titration to pH 4.5 in which 50 mL of FPW was titrated with 0.1 M HCl using a Metrohm Titrando 905. The milliequivalents (meq) of acid required to lower the pH was used to calculate the total alkalinity of the sample as ppm of CO_3^{2-}.

2.4.5 Iron (II) Analysis

The concentration of Fe(II) was determined using the ferrozine assay as described in Porsch and Kappler[53]. In summary, 400 µL of 1 M HCl was added to 100 µL of each sample, followed by an additional 500 µL of 0.1% by weight ferrozine. Samples were allowed to react for 5 min in the dark before spectrophotometric analysis at 562 nm using an Evolution 60S UV-Vis Spectrophotometer (Thermo Scientific). External calibration standards were prepared using Fe(II) ammonium sulfate to cover a range of 0.05-55 ppm Fe(II). Samples were diluted as needed to be within the calibration range.

2.5 Microscopy

2.5.1 Scanning Electron Microscopy (SEM)

A portion of the solids from each sample collected for digestion were set aside for imaging. Samples were then added to SEM tabs using carbon tape by dabbing the tab and tape onto the

196 filters. Imaging was performed using a Zeiss Sigma 300 VP-FESEM in variable pressure mode
197 with energy dispersive X-ray spectroscopy (EDS) with dual silicon detectors, each with 60 mm^2
198 of surface area and having a resolution of 123 eV. For better resolution of particle morphologies,
199 the samples were subsequently carbon coated using a Nanotek SEMprep 2 sputter coater before
200 reanalysis in high resolution mode under high vacuum.

201 **2.5.2 Transmission Electron Microscopy (TEM) with EDS and selected area electron**
202 **diffraction (SAED)**

203 The solids from S1 were suspended in 18MΩ ultrapure water at a concentration of approximately
204 1 mg/L. The suspension was then drop cast onto a 400 mesh carbon-coated copper grid (Pacific
205 Grid Tech), then air dried. Imaging was conducted using a FEI Titan 80-300 TEM operating at
206 80 and 200 kV with point-to-point resolution of 0.19 nm and a single tilt sample holder capable
207 of rotating from -30 to 30°. Elemental analysis was performed using an Oxford INCA 30 mm^2
208 LN_2 EDS detector with a 130 eV energy resolution, along with SAED.

209 **2.5.3 Scanning Transmission Electron Microscopy (STEM)**

210 A 1 g/L suspension of the solids from S1 was placed on ultra-thin carbon-coated copper grids.
211 Excess water was removed with paper filters and the residual was evaporated under a stream of
212 N_2 gas. TEM analysis was then conducted using a JEOL JEM-ARM200CF Atomic Resolution
213 S/TEM at the NanoFAB facility at the University of Alberta. Scanning transmission electron
214 microscopy (STEM) coupled to energy-dispersive X-ray spectroscopy (EDX) was done to
215 localize particles of interest and to generate elemental maps. Transmission electron microscopy
216 (TEM) and selected area electron diffraction (SAED) was then performed on selected particles
217 with a camera length of 60 cm. SAED patterns were analyzed following the manufacturer's

instructions. Specifically, the distance (in pixel) of diffraction peaks to the center beam and d-spacing was measured using the Gatan Microscopy Suite Software.

3 Results and Discussion

3.1 Aqueous Chemistry

The TDS was found to range from 171,000 to 243,000 ppm, with the primarily solutes for all three samples comprising Na and Cl (85-97%). While the TDS is known to generally increases with time after FPW has started[19, 20, 22, 23], the TDS of S3 was composed of the highest percentage of Na and Cl (97%) and was within 6% of S2 (91%) even though S2 was collected 9 days later. This indicates that variables, such as shut-in time and the use of recycled FPW for the mixing of HFF, effects the temporal variation of FPW's composition. The higher fractions of Na and Cl in sample S3 was likely due to approximately 10% of the makeup fluid for the HFF used to fracture well S3 being recycled FPW from other wells. Other major elements (those greater than 0.5%) were Ca, K, and Sr, descending order of their relative concentration. Trace elements (those less than 0.5%) Br, Li, Mg, Mn, B, Zn, Ba, and S were detected in every sample (Table 1). Aqueous Si was only detected in S1, which is not unexpected given the low solubility of Si minerals at room temperature, including amorphous silica[54,55]. The presence of Ba in the FPW is likely related to the dissolution of barite or desorption from aluminosilicate clays as has been suggested by Renock et al.[56] into the fracturing fluids. Sr in the FPW is likely linked to a combination of inclusions in carbonates and barite, in addition to desorption from clays much like Ba. Further investigation and leaching studies would be required to full understand the distributions of Ba and Sr and their subsequent mobilization. Silica could come from a variety of sources including the dissolution of quartz in the formation or the sand used as proppants. The elevated subsurface

241 temperatures are likely not great enough on their own to explain the aqueous Si concentrations,
242 and likely it, in combination with the complexation of Si with organic matter, plays a critical role
243 in the dissolution of the quartz as has been observed in room temperature systems[57,58].

244 The total organic carbon (TOC) concentration was the highest in S2 at 737 ppm, which
245 was more than double S3 at 294 ppm and S1 at 212 ppm. As observed by He et al.[1] for S1, the
246 total nitrogen (TN) was higher than TOC for S3 at 410 ppm; however, this was not observed in
247 S2 at 425 ppm (Table 1). The relatively lower TOC than TN is likely due to most organic
248 compounds in the samples being associated with an immiscible phase that floated on top of the
249 aqueous samples, and which was, therefore, not collected by the autosampler during analysis[1].

250 The presence and persistence of aqueous Fe(II) found in S3 (254 ppm) and S1 (35.7 ppm)
251 indicates that either weak or transient oxidizing conditions exist in the subsurface through the
252 introduction of HFF. It is unclear whether all the Fe(II) mobilized during fracturing returns to the
253 surface or only a portion that is the residual remains after the precipitation of Fe(III) minerals
254 that result from the subsurface oxidation of the mobilized Fe(II), as suggested by Jew et al.[59].
255 The fact that Fe(II) has been found in the FPW suggests that the reducing conditions were
256 reestablished in the subsurface post fracturing. The most likely source of Fe is pyrite as
257 experiments with Marcellus shale materials under hydraulic fracturing temperatures and
258 pressures have found a correlation between the mobilization of Fe and the initial pyrite content
259 of the shale used in the experiment[49]. Additionally, leaching experiments on Horn River
260 Formation shales (British Columbia, Canada) with DI water in equilibrium with atmospheric
261 oxygen found the oxidation of pyrite released sulfate and Fe(II)[60]. Elevated levels of Fe,
262 presumably Fe(II), is a widespread observation in the FPW of numerous formations in addition
263 to the Duvernay Formation, including the Marcellus Formation (0.3-747 mg/L)[20,21,61], Barnett

Formation (50-300 ppm)[71], Bakken Formation (0.7-96 ppm)[62] and Bowland Formation (36.6-137 mg/L)[64]. The aeration of S3 resulted in a 75.6% decrease in aqueous Fe(II), and in terms of color, from transparent yellow to turbid orange (Figure SI.2). Collectively, the decrease in aqueous Fe(II), change in color and increase in turbidity are consistent with the precipitation Fe(III) oxyhydroxides, such as ferrihydrite ($Fe(OH)_3$), as has been previously hypothesizes to account for losses of aqueous Fe during the handling of Marcellus FPW by Phan et al.[38] and Tasker et al.[65].

3.2 XRD Analysis

Both samples of drill cuttings from well 3 were found to contain quartz, calcite, muscovite, and aragonite (Figure SI.3). Sample A was found to contain sanidine, while sample B had kaolinite. The results of the bulk analysis are in generally in agreement with the literature for the composition of Duvernay Formation. The shales and argillites of the Duvernay have been shown to be primarily composed of quartz, mica, feldspar, plagioclase, calcite, and illite/smectite with less amounts of kaolinite[47,48]. There due to the heterogeneity within the formation it is not unsurprising that pyrite and barite might not found in the two samples which can be locally common[48,66].

3.3 Solids Analysis

The chemical composition of the solids for each sample were dominated by Fe and Si, with lesser amounts of S, Ca, Sr, and Ba (Figure 1 and SI.4 for mg/kg). Trace amounts (less than 100 mM) of Al, K, Mg, P, Mn, and Zn were detected, however, not in every sample (Table SI.1). The largest variations between samples was their ratios of Si to Fe, not their elemental composition. S1 was predominately Fe-rich, with the next most abundant element being Si (3351 mM/kg); S,

Ba, Ca, and Sr were the only other elements greater than 99 mM/kg. In contrast, S2 had an excess of Si to Fe 5842 to 1630 mM/kg respectively. Additionally, S2 had the highest amounts of S (778 mM/kg), and Sr (717 mM/kg). S3 also had a Si:Fe ratio greater than 1 (1.34) and had the highest ratios of Ca and Ba, 0.21 and 0.16, respectively. The Sr+Ba:S molar ratio for all samples was approximately 1 and 0.89, 0.98, and 0.94 for S1, S2, and S3, respectively. In contrast, the relative abundances of Ba and Sr do vary: S1 and S3 have a relative excess of Ba to Sr, 1.37 and 1.42, while S3 contained a low Ba to Sr ratio, 0.07.

Aerating S3 (S3_Ox) resulted in the decline of the relative ratio of Si/Fe from an excess of Si to an excess of Fe (Si/Fe = 0.41). The ratios of Si, Ba, Ca, and Sr to Fe also declined. Aeration was effective in the removal of dissolved Fe(II) and the concomitant generation of Fe(III) precipitates. The Ba/S ratio between samples S3 and S3_Ox were similar at 0.55 and 0.59, respectively, while the ratio of Sr/S showed a Sr enrichment in S3_Ox (0.61) compared to S3 (0.39). The removal of Sr during aeration is also evident in the relative abundance of Ba/Sr which shifted from 1.41 (excess Ba) to 0.98. The removal of Sr from solution during aeration suggests its adsorption or incorporation during the precipitation of the Fe(III) oxyhydroxides. The incorporation of Sr into iron oxides has previously been observed during ferrihydrite ripening to goethite[67].

3.4 Morphological and Chemical Characterization

The backscatter SEM images (Figures 2 A, C, E, and G) show the solids had two particle morphologies: (1) globular neoformed aggregates, and (2) stellated clusters For both particle types, the strong C signals observed were attributed to the carbon coating used to prepare the samples and the carbon-based tape used to affix samples to the aluminum tabs, and the strong O

signal primarily to O associated with the mineral oxides that comprise the solids. We cannot exclude the possibility that a portion of the C and O signal was related to organic compounds associated with the sample, however, with the carbon tape it is impossible to quantify what portion that may be. Investigating whether there is significant insoluble organic carbon associated with the FPW solids will need to be investigated in the future. For the purposes of this study, C and O have been excluded from the discussion that follows. The EDS spectrum (gray bars) of S1's globular, neoformed aggregates have an approximate composition in molar percent of: Si (2.19%) and Fe (5.13%) (Figure 2 B). The neoformed aggregates in S2 and S3 showed similar elemental profiles (Figure 2 D, H). However, for S3_Ox, an increase in the relative molar percent of Fe was observed (Figure 2F and H). While the ratio of Fe to Si varied between samples and EDS spots, neither element was found independent of the other. These observations suggest that the neoformed aggregates are discrete Fe/Si phases, representative of a silica-enriched Fe(III) oxyhydroxide.

Representative EDS spectra targeting the stellated clusters in S1 (Figure 2 B) show that they are composed of S (4.19%), Sr (1.54%), Ba (3.02%), Si (1.54%), and Fe (2.78%), with a close association between S, Ba and Sr. stellated clusters from S2, S3, and S3_Ox exhibited similar elemental profiles (Figures 2 D, F, and H). This correlation was not observed in the neoformed aggregates. Fe and Si were not observed in every spectrum with S, Ba, and Sr. We attributed the low presence of Fe and Si in the EDS spectra as an artifact of the large EDS spot size, which resulted in the EDS beam also contacting part of the neoformed aggregates silica-enriched Fe(III) oxyhydroxide. The consistent elemental profile of stellated clusters indicates that they are singular phase composed of Sr, Ba, and S, with a Sr+Ba:S molar ratio of approximately 1. This is consistent with a sulfate mineral such as barite or celestine.

Across all three samples, the neoformed aggregates of the untreated samples showed a similar spherical particle morphology and size (~200 nm) (Figure 3 A, B, and C). Aerating S3 generated additional neoformed aggregates of a similar size and morphology to those found in the untreated samples (Figure 3 D). A comparison of S3 and S3_Ox revealed that S3 had plates of granular aggregates composed of 200 nm spherical particles with a few aggregates associated with the surfaces, while S3_Ox has numerous colloidal aggregates on the surfaces (Figure 3 A and B).

HRTEM with EDS and SAED was further used to probe the composition and minerology the Fe/Si phase. HRTEM images of S1 reveal a lack of order amongst the neoformed aggregates, with no visible edge steps (Figure 4 A and B). The elemental composition was dominated by Fe and Si (Figure 4 D). The SAED pattern generated from the Fe/Si aggregates (Figure 4 C) showed no apparent diffraction pattern, indicating a lack of crystallinity consistent with an amorphous silica enriched Fe(III) oxyhydroxide phase. STEM elemental maps of the solids from sample 1 show a strong association between Fe, Si and O (Figure SI.5 A, B, and C). The Si appears to have a higher abundance at the core of the particle with the Fe aggregating on the Si core. The elemental map of the Sr+Ba:S phase show that a higher concentration of S compared to Sr or Ba and that all three elements are concentrated in the particle (Figure SI.6 A, B, C, and D)

3.5 Modeled Saturation Indices

The saturation indices of each sample were calculated as a function of temperature, from 10 to 120° C. This represents the approximate range of temperatures that FPW would encounter as it returns from the elevated subsurface temperatures, approximately 115° C for the Duvernay

Formation[68], and equilibrates at the surface. The thermodynamic modeling was performed in PHREEQC[69] utilizing the Pitzer database due to the high ionic strength of the FPW samples 4.11, 3.24 and 3.31 M for S1, S2, and S3, respectively. The Pitzer database included barite ($BaSO_4$), celestine ($SrSO_4$), quartz (SiO_2), and amorphous silica ($SiO_{2(am)}$). To model the SI, the measured concentrations from S1, S2, and S3 of all elements found to be greater than 100 ppm (Na, Cl, Ca, K, Sr, Br, and Mg) and Fe, S, Ba, and Si were utilized, with the pH set at that measured for each respective sample. For the purpose of this model all S from the total dissolved S measurement was assumed to be sulfate (SO_4^{2-}) during modeling. S speciation in flowback is known to include thiosulfate and hydrogen sulfide, however, they have been found to be relatively minor components totaling less than 10% the sulfate concentration[70]. To begin, the sensitivity of the saturation indices to changes in pH was examined the models were run as pH sweeps from 3-10. The saturation indices for barite and celestine were found to not be sensitive to pH across the pH range examined (Figure SI.7 A and B), while the saturation indices of silica was found to decrease exponentially above pH 8, well above the pH of the FPW samples examined (Figure SI.8). Additionally, the models were examined for to determine whether the exolution of CO_2 may have occurred during the cooling process, which could have influenced changes in FPW pH, however, across the temperature range examined the saturation indices for CO_2 was well below 0 (Figure SI.9).

S1, S2, and S3 were found to be supersaturated with respect to barite at room temperature (20°C). The saturation indices show a distinct trend of increasing with decreasing temperature, which indicates that it is more thermodynamically favorable to form a solid phase as the solution cools (Figure 5 A). S3 had the highest equilibrium temperature, approximately 90°C, which may be attributed to its being the freshest sample and the process of precipitation being kinetically

slow. In contrast, all three samples were found to be undersaturated with respect to celestine with saturation indices ranging from -0.28 to -0.66 across the entire temperature range examined (Figure 5 B). As the models were performed using elemental analysis collected on solutions that had been fully cooled, this does not mean that the solution was not previously supersaturated with respect to celestine. It is more plausible that the Sr concentration was initially drawn down during the cooling by the precipitation of celestine. The continued precipitation of barite after the celestine reached equilibrium likely drew down the sulfate concentration and potentially the Sr concentration through incorporation to reflect the measured undersaturated concentrations. Interestingly, S2 had the lowest saturation indices and the lowest Ba/Sr ratio in the sediment (0.070).

While Si was found in solids of all three samples, Si was only above the detection limit in the aqueous phase of S1. For S1 we modelled the saturation indices for this sample with respect to the two silicate minerals quartz and amorphous silica. At room temperature, S1 was supersaturated with respect to quartz but not $SiO_{2(am)}$. S1 was supersaturated with respect to quartz below 60°C; in contrast the lowest saturation indices for $SiO_{2(am)}$ was -0.67 (Figure 6). Furthermore, we modeled each sample in equilibrium with respect to either barite and celestine or quartz and $SiO_{2(am)}$ to determine how the measured Si, Ba, Sr, and sulfate concentrations compared to the modeled equilibrium concentrations. Regarding the silicate minerals, S1 had an excess of Si with respect to quartz below 80°C, respectively. In contrast, S1 was below the modeled Si equilibrium concentration with respect $SiO_{2(am)}$ across the entire temperature range, but by less than 1 mM when the temperature was below 30 °C (Figure SI.11). Even if we assume samples S2 and S3 had no measurable Si in solution, they too are still only below the equilibrium

concentration of Si by less than one mM for quartz across the entire temperature range, and by less than one mM when the temperature is at or below 20 ℃ for $SiO_{2(am)}$.

For S1, the system was found to be supersaturated with respect to sulfate across the modeled temperature range and to have an excess of Ba when the temperature was below 50℃. Across the examined temperature range, Sr was below the equilibrium concentration by 2.26-3.04 mM (Figure SI.10 A). Similarly, S3 had an excess of sulfate and Ba across the entire temperature range and Sr was below the equilibrium concentration by 2.62 - 3.59 mM (Figure SI.10 C). In contrast, S2 was below the equilibrium concentrations of sulfate and Sr across the entire temperature range, by 56.4 and 20.1 mM respectively, while below 70° C there was an excess of Ba (Figure SI.10 B).

3.6 Multi-basin Comparison

Publicly available FPW chemistry data for the Fayetteville (4 samples)[70], Marcellus (day 14 flowback median from 7 wells, average of 95 well, 2 composite samples of days)[20,21,34], Bakken (median of 4 wells from day 10-12)[62], Barnett (median from day 10-12 from 4 wells)[63] and Denver-Julesburg (DJ) Formations (4 wells)[73] were modeled for using the same PHREEQC model developed for the Duvernay Formation (Full modeling parameters are shown in Table saturation indices 2). While each basin is expected to have heterogeneity, these models offer a first pass determine how similar the FPW solids are between shale basins and how widely applicable treatment options would be. In the case of absent pH data for the Fayetteville and Marcellus, the pH was assumed to be 7 as the models only showed a pH dependence above pH 8.

For barite, below 90^0 C the precipitation was found to be favorable (saturation indices greater than 1) for FPW from the Marcellus, Bakken, Barnett and DJ formations for all sample

where the sulfate concentration was reported (Figure 7 A). Below 30^0 C this also included the Duvernay. The Fayetteville was the only formation for which the precipitation of barite was not thermodynamically favorable across the temperature range. In contrast the precipitation of celestine was favorable for 3 out of the 4 Marcellus samples across the entire temperature range examined (Figure 7 B). For all other basins, the formation of celestine was not favorable at any temperature. However, as observed in this study for the Duvernay, this does not preclude the possibility of Sr incorporation into barite or the formation of a solid solution of Sr/Ba(SO_4). For all basins save the Fayetteville the saturation indices for celestine was greater than -1, meaning it is close to being in equilibrium with celestine.

Saturation indices for quartz and $SiO_{2(am)}$ was only modeled, in addition to the Duvernay Formation, for Fayetteville and DJ as dissolved silica data was not available for the Bakken, Marcellus and Barnett. $SiO_{2(am)}$ was found to have a saturation indices greater than 0 only for 1 out of 5 of the Fayetteville FPW samples (Figure 8 A). In contrast the formation of quartz was found to be favorable in all three basins with dissolved Si data below 50^0 C (Figure 8 B). Additionally, for the Marcellus while the absence of dissolved silica data precluded the modeling of saturation indices silica has been identified through SEM EDS analysis of solids fouling filters used for Marcellus FPW[34]. This indicates the absence of dissolved silica data or when it is found to be below the detection limit would not mean that the formation of quartz would not be favorable.

Additionally, appreciable concentrations dissolved Fe were reported in each of the basins Fayetteville (1-13 ppm)[71], Marcellus (40.8-747 ppm)[20,21,34], Bakken (0.7-96 ppm)[62], Barnett 12.1-93.8 ppm)[72] and DJ (2.71-19)[73]. This indicates that the precipitation of insoluble Fe(III)

oxyhydroxides is likely across more than the Duvernay and Marcellus Formations where it has already been observed[38,65].

3.7 Mechanisms of Formation

The amorphous Fe/Si phase likely formed through a two-step process. First, the FPW returned to the surface from the target formation with an elevated temperature (~115°C for the Duvernay[68], 82°C for the Fayetteville[74], 104 °C for the Marcellus[75], 71-116 °C for the Bakken[76], 93 °C for the Barnett[74], and 116 °C for the DJ Formation[77] and enriched in dissolved Si relative to surface temperatures[54,55]. As the FPW cooled, Si became increasing supersaturated with respect to amorphous silica. Second, with the increasing oxygen solubility and diffusion associated with cooling, the dissolved Fe(II) oxidized to Fe(III), and through subsequent hydrolysis formed insoluble Fe(III) oxyhydroxides. The Fe(II), and at least a portion of the sulfate, was probably derived from the oxidation of pyrite (FeS_2) by oxygenated fracturing fluid. In this process sulfide in pyrite in the formation oxidized to sulfate, consuming the oxygen in the fracturing fluid releasing sulfate and Fe(II) into solution[60]. The subsurface conditions rapidly return to anoxic after the oxygen is consumed through sulfide oxidation, which prevents the Fe(II) from also oxidizing to Fe(III). The Fe(III) oxyhydroxides are always closely associated with amorphous silica, an unsurprising observation given that Fe(III) oxyhydroxides have a high sorption capacity for dissolved silica[78,79]. However, in our study the evidence points towards the adsorption and templating of Fe(III) oxyhydroxides onto amorphous silica to be the process by which this phase formed. Additionally, the diffusion of oxygen may be limited by thin surface oil films on the FPW, as were observed in our samples. Nonetheless, the residual Fe(II) can be oxidized through treatments, such as aeration (see S3_Ox).

The mechanism(s) by which the FPW solids form similarly sized Fe/Si particles (~200 nm diameter) could be related to several factors including, the rate of cooling and/or Fe(II) oxidation, high ionic strength, or presence of organic matter (OM). OM is known to alter Fe(II) oxidation rates, control the final mineral formed during mineralization, and in tandem with Si, prevent the ripening of low order iron-oxide species into crystalline forms[80, 81]. The incorporation and adsorption of OM by iron-oxides can alter their surface charge preventing aggregation or agglomeration and enhancing their incorporation of contaminants[82,83].

The second solid phase, a mixed Sr/Ba sulfate, is representative of a barite-celestine solid solution. $BaSO_4$ and $SrSO_4$ have been found in Marcellus FPW[84] and their solubilities are temperature dependent[85,86]. The Sr/Ba-SO_4 phase likely formed as a function of the FPW cooling from the subsurface to surface temperatures and creating supersaturated conditions that drive their precipitation. Its precipitation is likely rapid, as barite scale formation increases with increasing ionic strength with Na initiating the growth of new step rows; the latter is the rate limiting step in low ionic strength solutions[87].

3.8 Environmental Implications

The characterization of FPW solids and will enhance the development of new FPW treatment options for reuse, contaminant removal, and transportation risk reductions. To date, the solids have been observed to account for approximately 50% of FPW's PAH load[1,45] and approximately 50% of their total toxicity as measured by LC_{50} and EROD[1,46]. Other organic contaminants may also be associated with the solids as only targeted analysis of PAHs have thus far been conducted. As Fe(III) oxyhydroxides were the dominant solid phase, and they are known to adsorb organic contaminants such as PAHs[88], we hypothesize that majority of PAHs

associated with the solids are sorbed to the Fe(III) oxyhydroxides. Furthermore, the fact that the aeration of the FPW generated additional approximately 200 nm diameter Fe/Si particles opposed to resulting in the growth of the previously existing particles is important. These freshly generated Fe/Si particles represent a large increase the particle reactive surface area in the FPW per litre and should enhance the ability of the solids to remove contaminants. Additionally, during simulated hydraulic fracturing experiments with shales containing pyrite and simulated hydraulic fracturing fluids containing dissolved oxygen, Xiong et al.[49] observed the generation of hydroxyl radicals and degradation of one HFF chemical, polyacrylamide. They hypothesized that the Fe(II) released by the pyrite in the presence of dissolved oxygen resulted in a free radical chain scission mechanism. Given the wide spread occurrence of presumably Fe(II) in FPW the aeration of FPW to promote the formation of Fe(III) oxyhydroxides by oxidizing residual Fe(II) has the potential to be a widespread treatment technique to remove additional organic contaminants, adsorb or incorporate heavy metals, and reduce the aqueous toxicity of FPW.

While NORMs have not appreciably been found in Duvernay Formation[27] FPW, they are a concern in other formations, such as the Marcellus Formation[25,26]. Ba and Sr sulfates are known to incorporate NORMs such as radium during their formation[35,89]. Therefore, NORMs and their associated risks could be concentrated in the solid phase. Adding sulfate to FPW has been proposed as a treatment to remove Sr and Ba from FPW[35]. Thus, understanding the sources of Ba[56] and removing it from FPW could promote the reuse of FPW by reducing the risk of barite scales[90]. The disposal of these solids, however, would need to account for any NORMs concentrated during treatment.

Additional work will be required to characterize the solids from other basin and understand the mechanisms by which heavy metals, NORMs, and organic contaminants, such as

PAHs[1,45,46], adsorb and desorb from FPW associated solids. This work could be important for remediating FPW spills. Furthermore, treatments that generate iron oxides could remove additional contaminants, reduce the environmental risks associated with FPW, and prevent their formation during the recycling of FPW. These insights will promote the further development of treatment technologies for the reuse of FPW.

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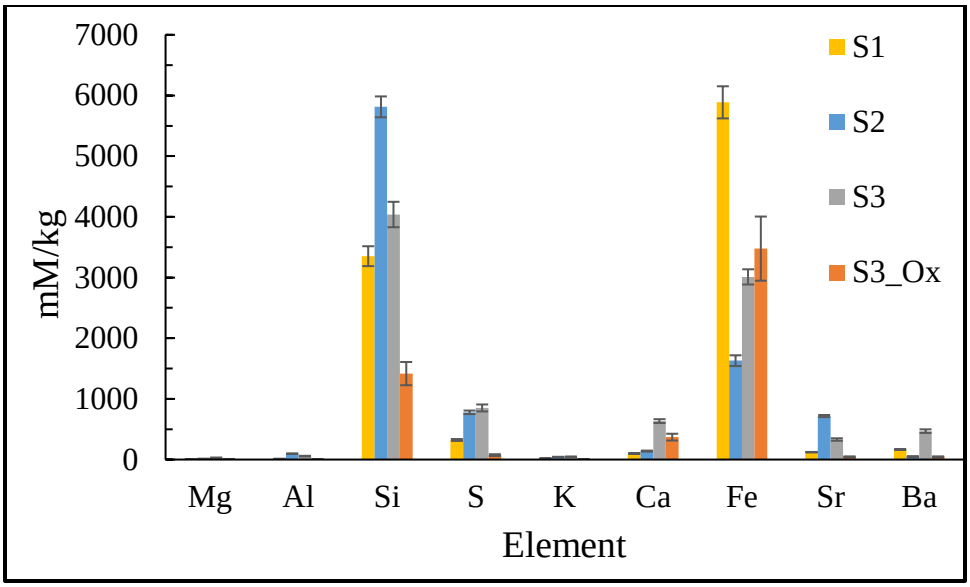
787

788 Table 1 Summary of the aqueous chemistry for the three samples analyzed in this study.
789 Reported concentrations and detection limits (DL) are in ppm unless otherwise noted. BDL
790 denotes below detection limit and NM indicates the parameter was not measured for the
791 corresponding sample. In brackets, the relative standard deviation for each analysis is presented
792 as the percent within one deviation.

	DL	S1*	S2#	S3
TDS		243,000	183,000	171,000
pH		4.78	5.86	5.51
Alkalinity (as CaCO3)	NM	2.25	14.3	54.4
T (C°) at collection		60	NM	38
TOC	0.46	212	737	294
TN	0.27	498	425	410
Cl	0.04	136,000	107,000	112,000
Na	0.33	70,000	59,500	54,300
Ca	1.13	11,800	6,500	8,630
K	1.47	2,570	2,160	1,930
Sr	0.049	1,470	931	904
Br	0.24	276	297	242
Fe	0.23	22	1.19	151
S	0.18	64.5	66.9	78.3
Si	0.37	7.9	BDL	BDL
Ba	0.00022	5.85	7.28	24.1
SO_4^{2-}	0.2	48.7	BDL	BDL
Li	0.00038	54.6	51.3	41.3
B	0.021	71.6	96	79.1
Mg	0.090	111	706	738
Mn	0.00090	15.8	7.32	15.1
Ni	0.0055	BDL	BDL	BDL
Cu	0.01	BDL	BDL	BDL
Zn	0.018	4.4	1.24	0.561
As	0.0050	BDL	BDL	BDL
Cd	0.00013	BDL	BDL	BDL
Pb	0.00014	0.06	BDL	BDL

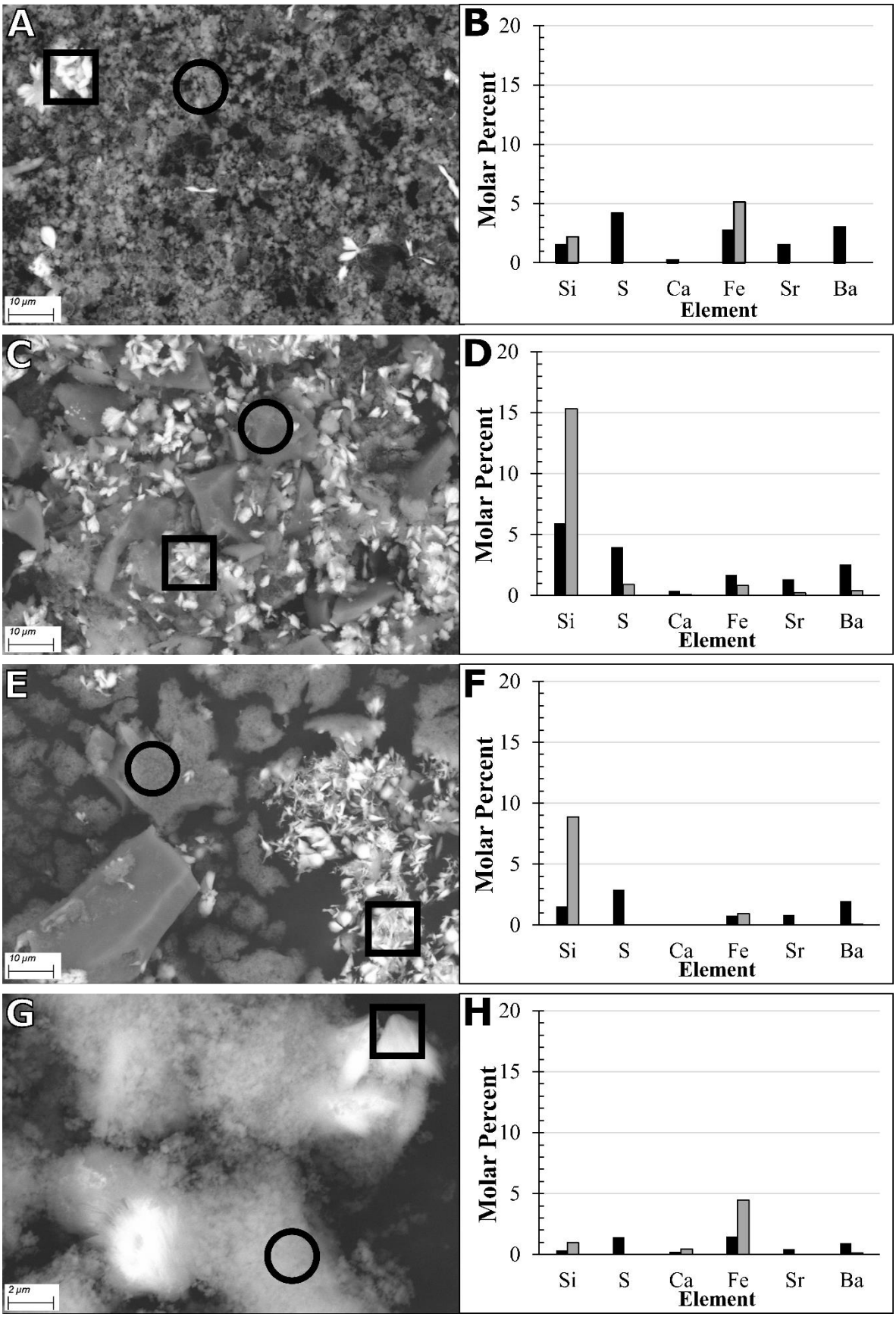
793

794 The chemistry of S1 and S2 were previously reported in *He et al.[1]; #Blewett et al.[50],
795 respectively, with the new analyses being conducted for alkalinity and total sulfur for both and
796 Pb, Cu, and Si for S2.

797 

798 Figure 1 Elemental profile of FPW solids as determined through alkaline fusion digestion, in

799 with the error bars representing ±1 standard deveation.


A
10 µm
B
Molar Percent
20
15
10
5
0
Si
S
Ca
Fe
Sr
Ba
Element
C
10 µm
D
Molar Percent
20
15
10
5
0
Si
S
Ca
Fe
Sr
Ba
Element
E
10 µm
F
Molar Percent
20
15
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Si
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Sr
Ba
Element

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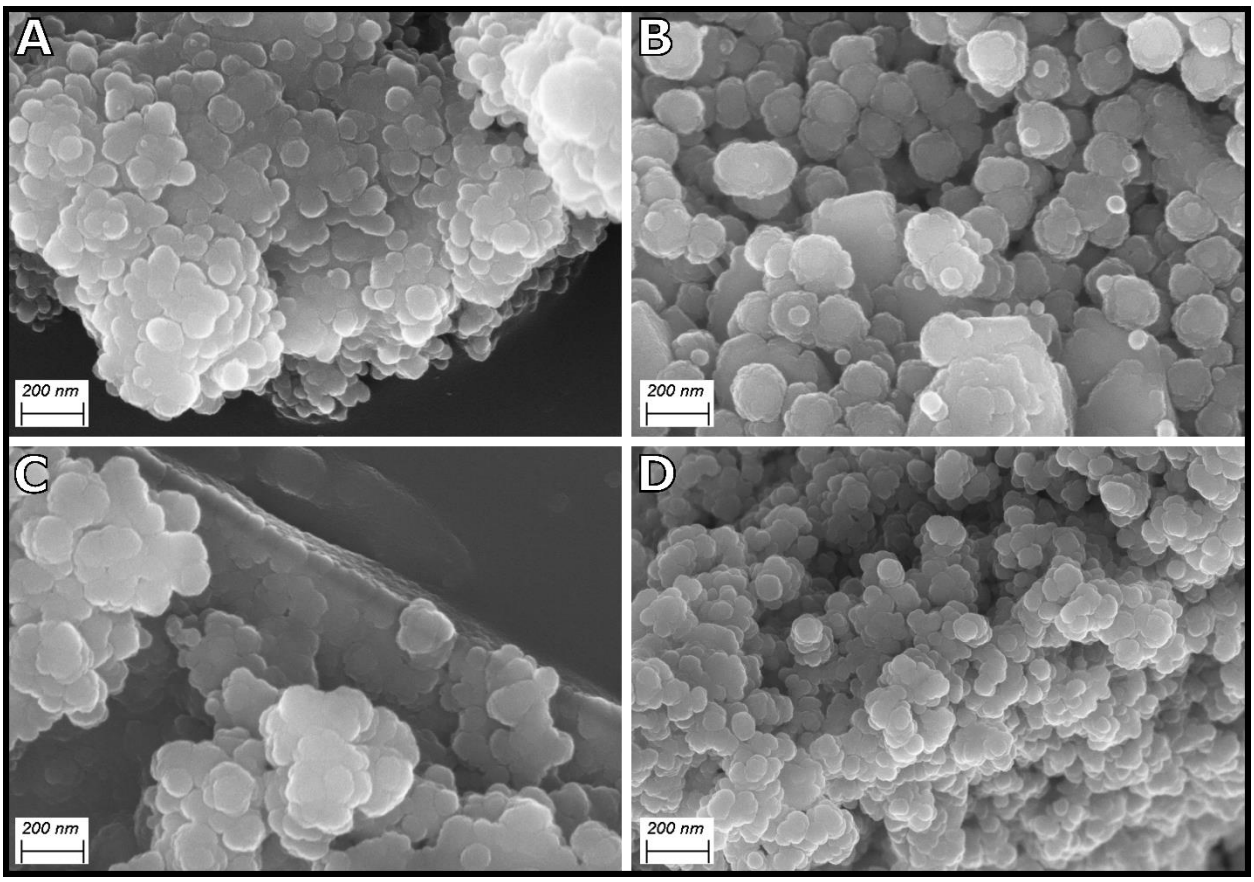
801 Figure 2. SEM images in backscatter mode and elemental distribution determined by EDS of
802 samples S1 A) and B), S2 C) and D), S3 E) and F), and S3_Ox G) and H). In each SEM image
803 the square and circle represent the location from which the EDS spectra was acquired and
804 correspond to the black and white bars in the corresponding elemental distributions as molar
805 percent, respectively.



806

807 Figure 3. High resolution SEM images of carbon coated samples showing the morphology of
808 Fe/Si globular neoformed aggregates associated solids: A) S1, B) S2, C) S3, and D) S3_Ox.

809

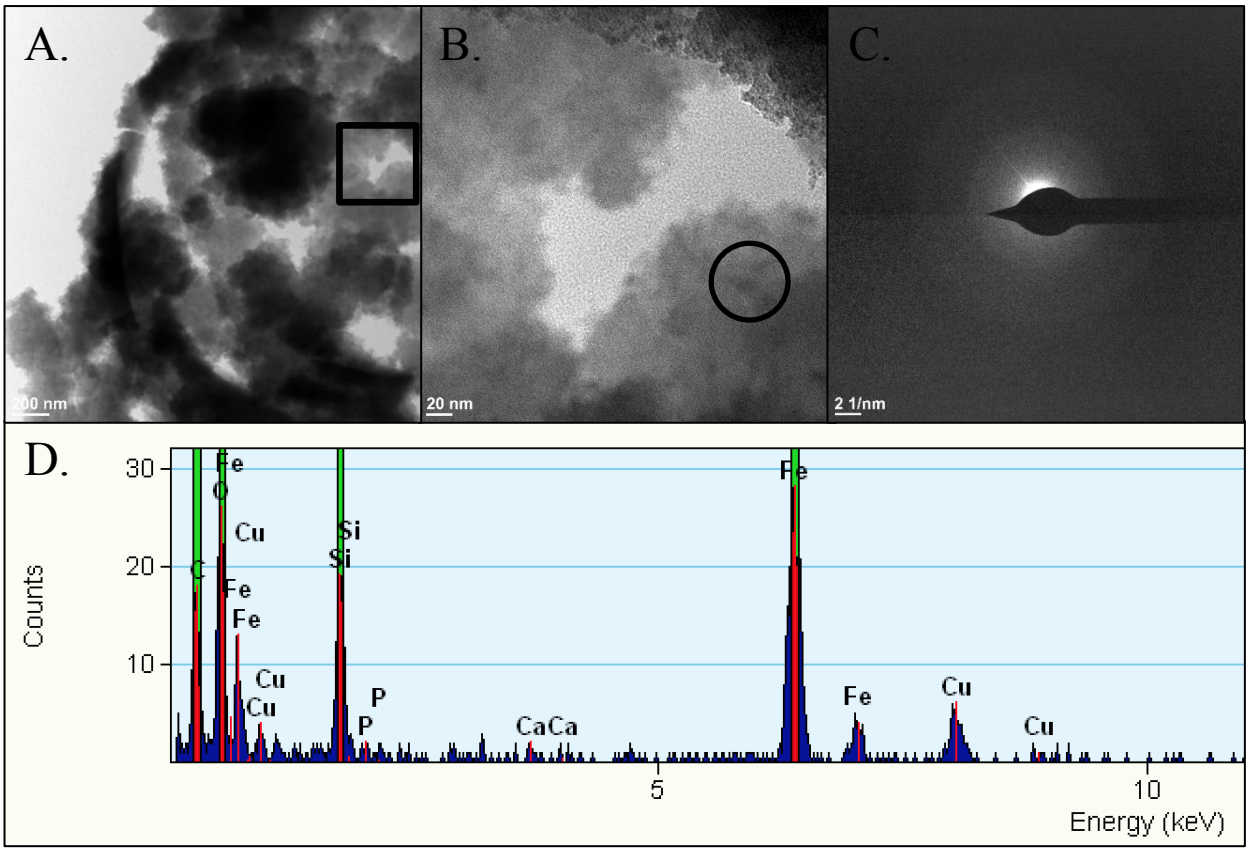


Figure 4. TEM analysis of solids of S1. A) A wide image in which the square represents the location of the close-up image in panel B). B) is a close-up image in which the circle represents the approximant location of the SAED spectra presented in panel C) and EDS spectra presented in panel D).

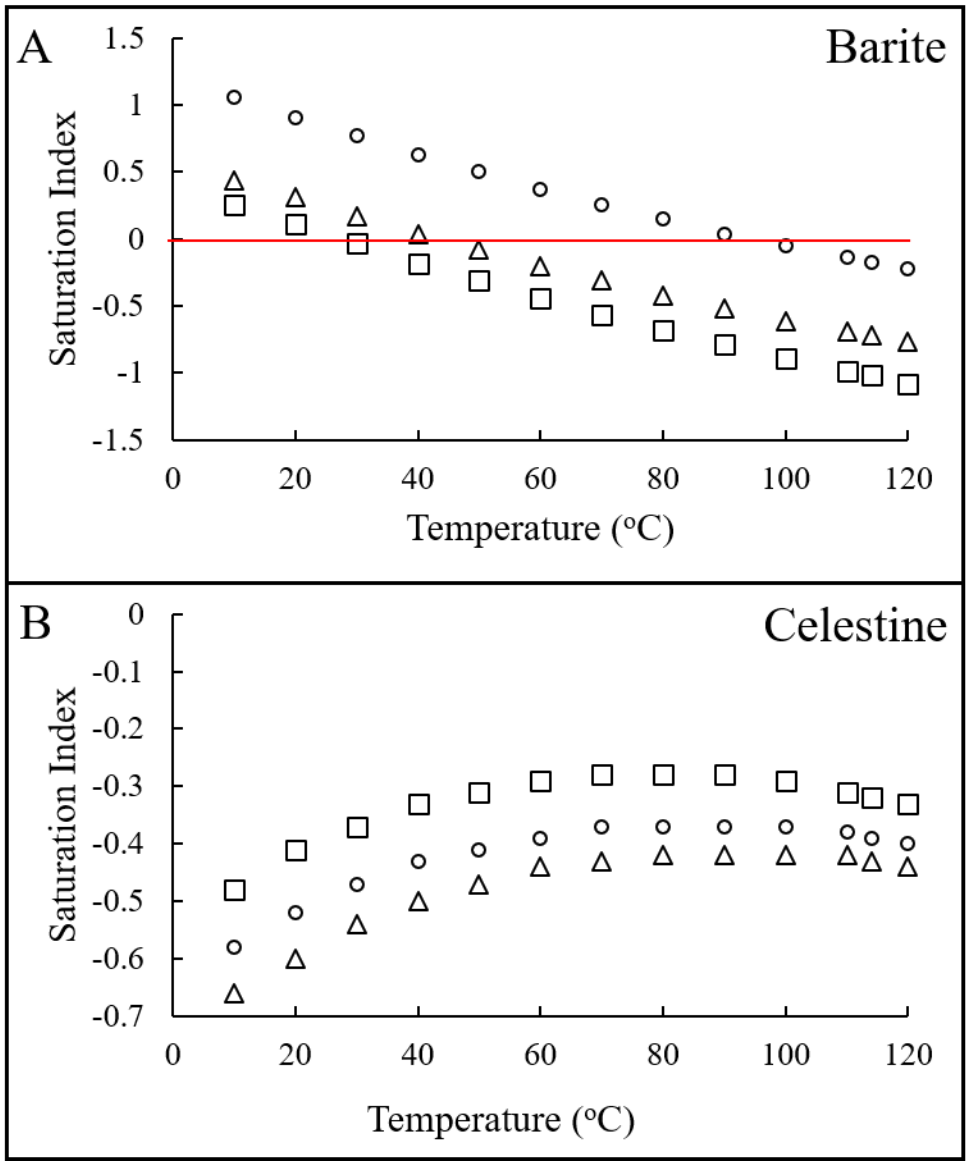
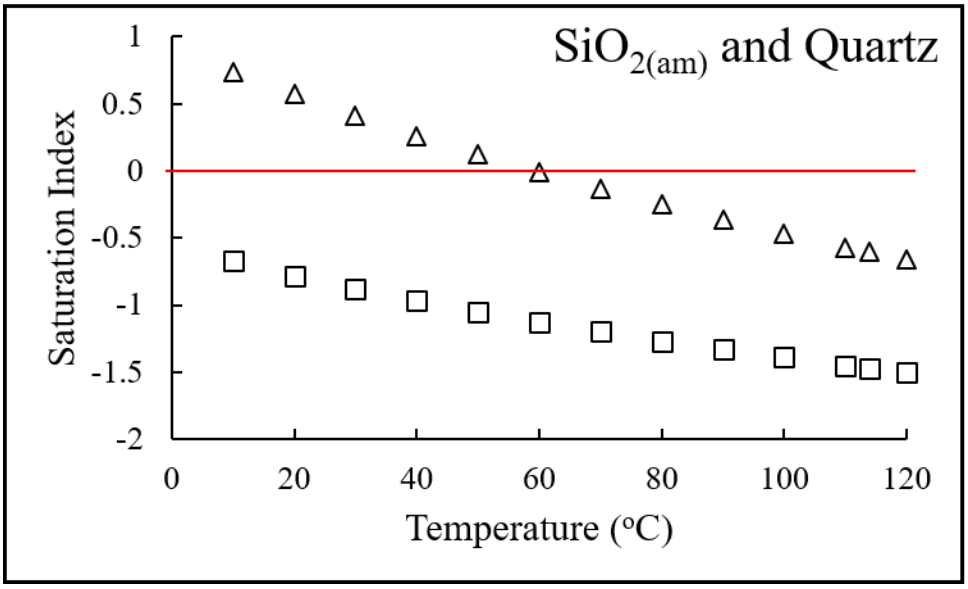
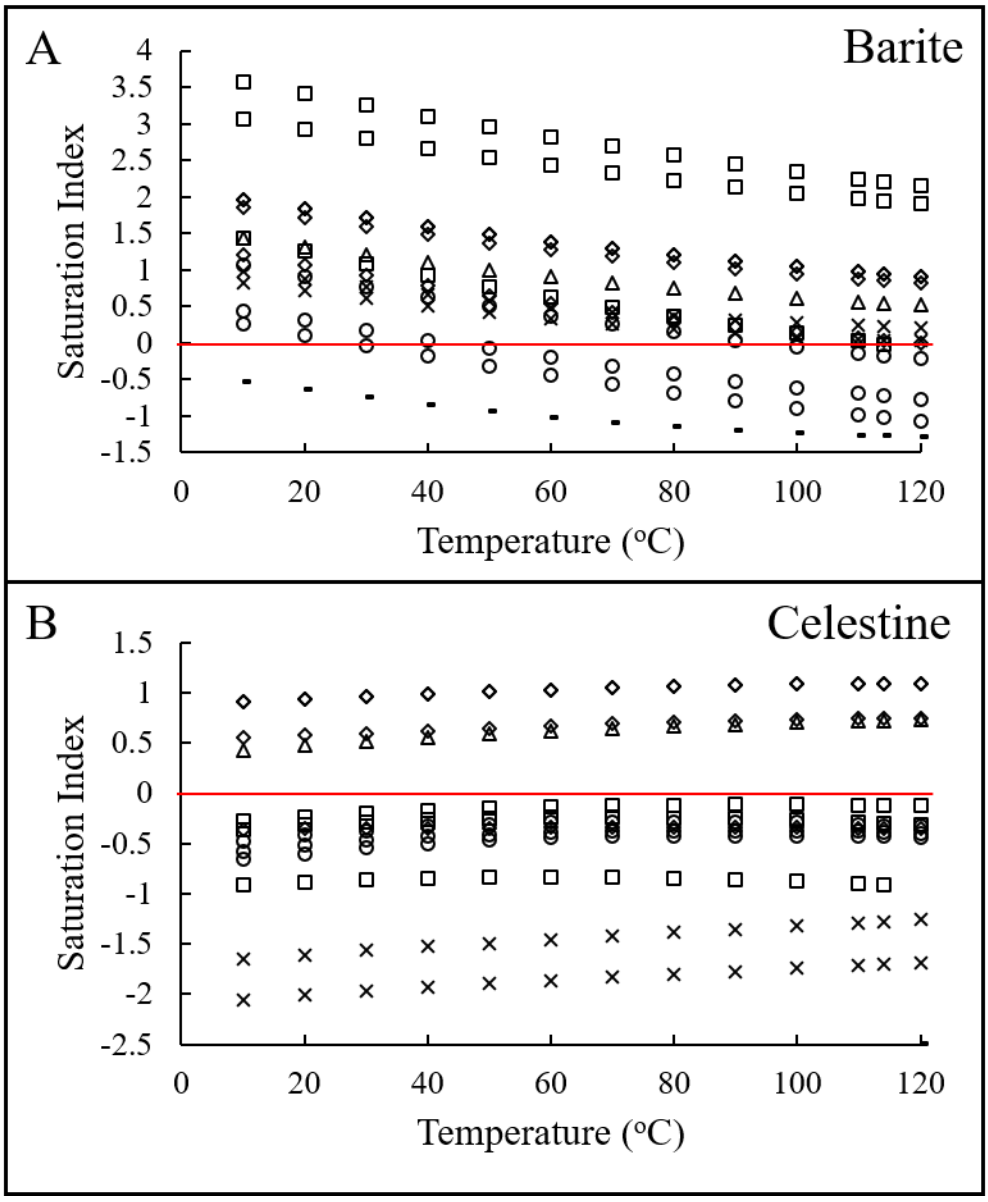


Figure 5. Saturation indices of hypothesized minerals present in Duvernay FPW modeled using the element data as a function of temperature A) respect to barite, and B) celestine, with S1, S2, and S3 represented by (□), (Δ), and (○), respectively.



827

828 Figure 6. Saturation indices of potential silicate minerals present in Duvernay FPW modeled
829 using the element data from S1 as a function cooling for $SiO_{2(am)}$ (□), and quartz (Δ).



830

831 Figure 7. A cross basin comparison of saturation indices of FPW for A) barite and B) celestine as
832 a function of temperature for the Duvernay, Marcellus, Barnett, Bakken, Denver-Julesburg, and
833 Fayetteville (○), (□), (Δ), (◊), (×) and (-), respectively.

834

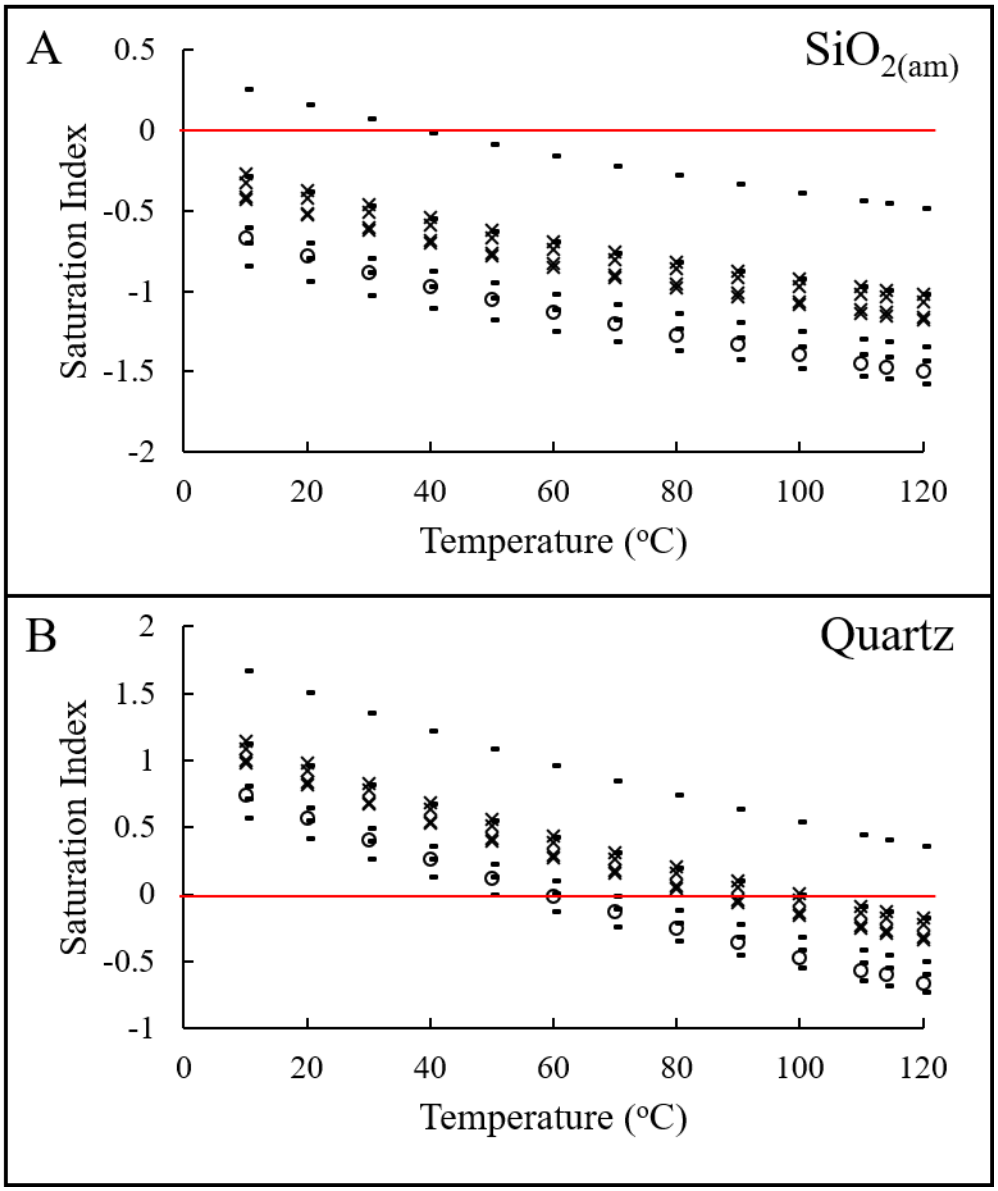


Figure 8. A cross basin comparison of saturation indices of FPW for A) $SiO_{2\,(am)}$ and B) quartz as a function of temperature for the Duvernay, Bakken, Denver-Julesburg, and Fayetteville (○), (Δ), (×) and (-), respectively.

842 **Supplementary Information**

843 Characterization and Implications of Solids Associated with Hydraulic Fracturing Flowback and
844 Produced Water from the Duvernay Formation, Alberta, Canada

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857

Duvernay Formation Geology and Composition

The Frasnian Duvernay Formation (Woodbend Group) is a bituminous and variably calcareous mudstone that is present throughout most of the Western Canada Sedimentary Basin. The Duvernay Formation is laterally equivalent to the Leduc Formation, a carbonate platform and reef complex that has been exploited for conventional oil and gas since 1947. The Geology of the Duvernay Formation is summarized in Stoakes[1]. The Duvernay Formation comprises Interbedded bituminous shales, dark brown, calcareous shales and dense argillaceous limestones. More rarely observed are fossil-rich lime mud accumulations. The bituminous shales are characteristically plane-parallel laminated. Argillites are also planar laminated, but may also be massive appearing or completely bioturbated. The variability in bedding is best interpreted as unsteadiness in oxygenation further suggesting that redox-sensitive elements are heterogeneously distributed. The shales and the argillites contain — in descending abundance — silt-sized quartz, mica, K-feldspar, and plagioclase; the clay-size fraction is dominantly quartz and calcite with subordinate clay illite/smectite and less abundant kaolinite[2,3]. Total organic content is variable, locally exceeding 10% and more commonly between 1 and 3%[4]. Eogenetic to diagenetic pyrite is locally common. Barite, likely from hydrothermal sources, is present up to 1.4 weight %[3].

Profile of Well 3

Well 3 was drilled into the Duvernay Formation at a depth of 3172 m with a horizontal bore length of 2535. Sample A came from 139 m into the horizontal bore while B came from 839 m (Figure SI.1)

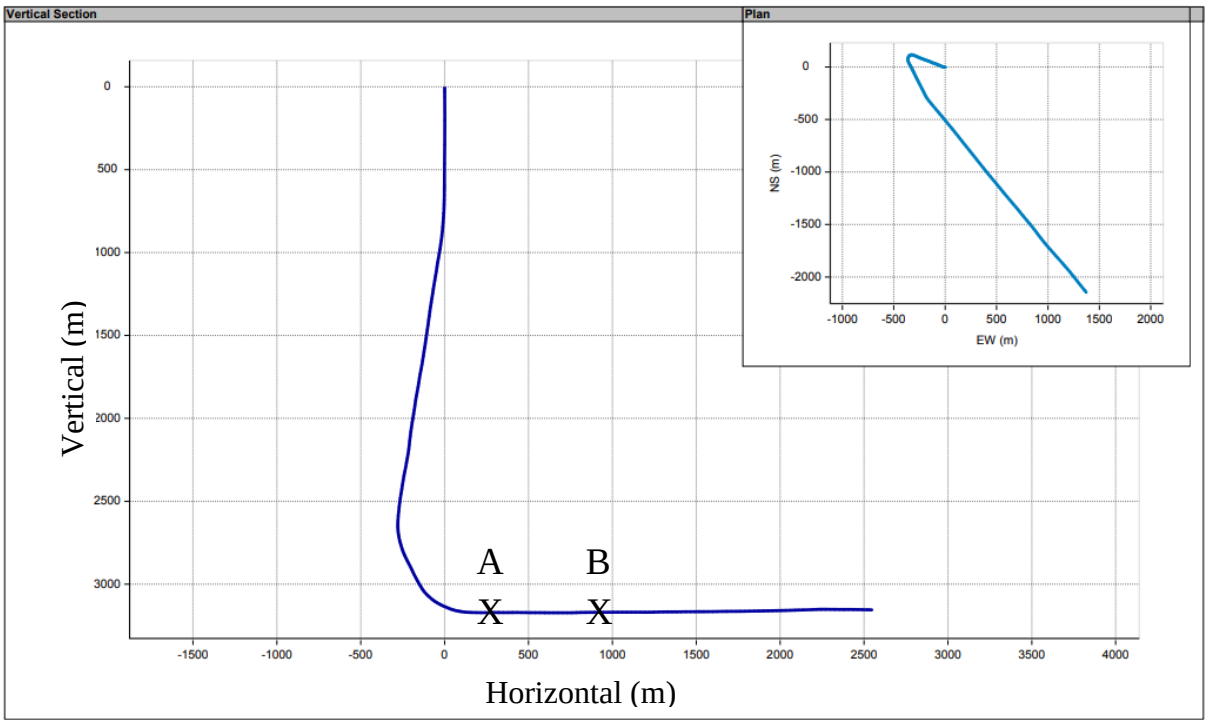
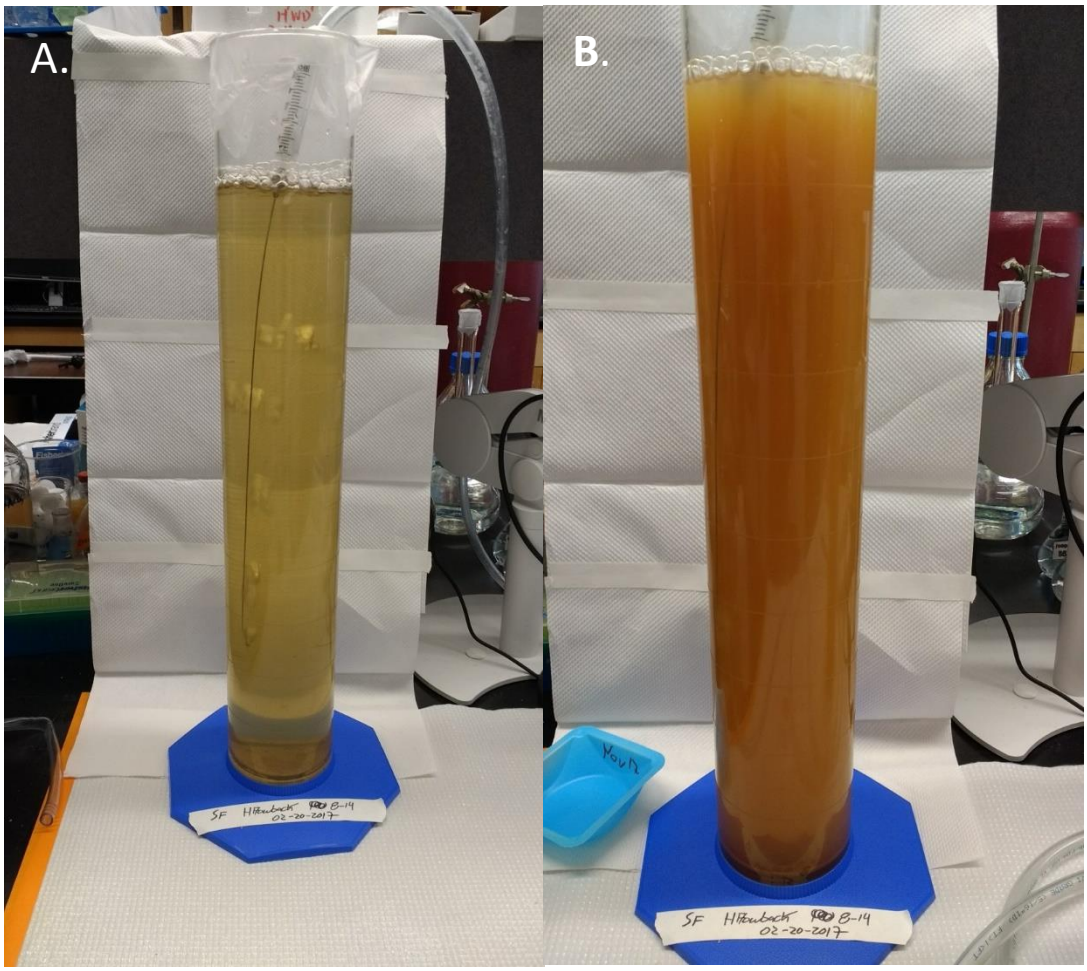
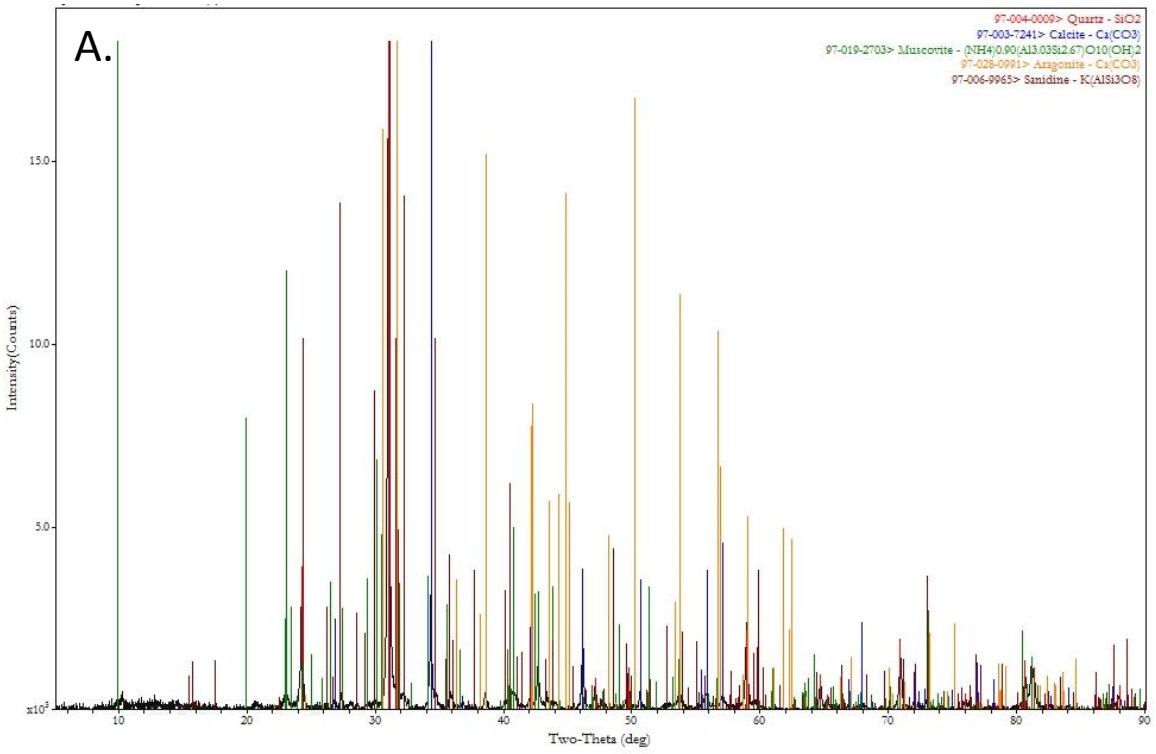


Figure SI.1 Schematic of well 3's bore showing the approximant locations from where drill cutting samples A and B were obtained.



883 Figure SI.2 Sample S3 before (A) and after bubbling (B).



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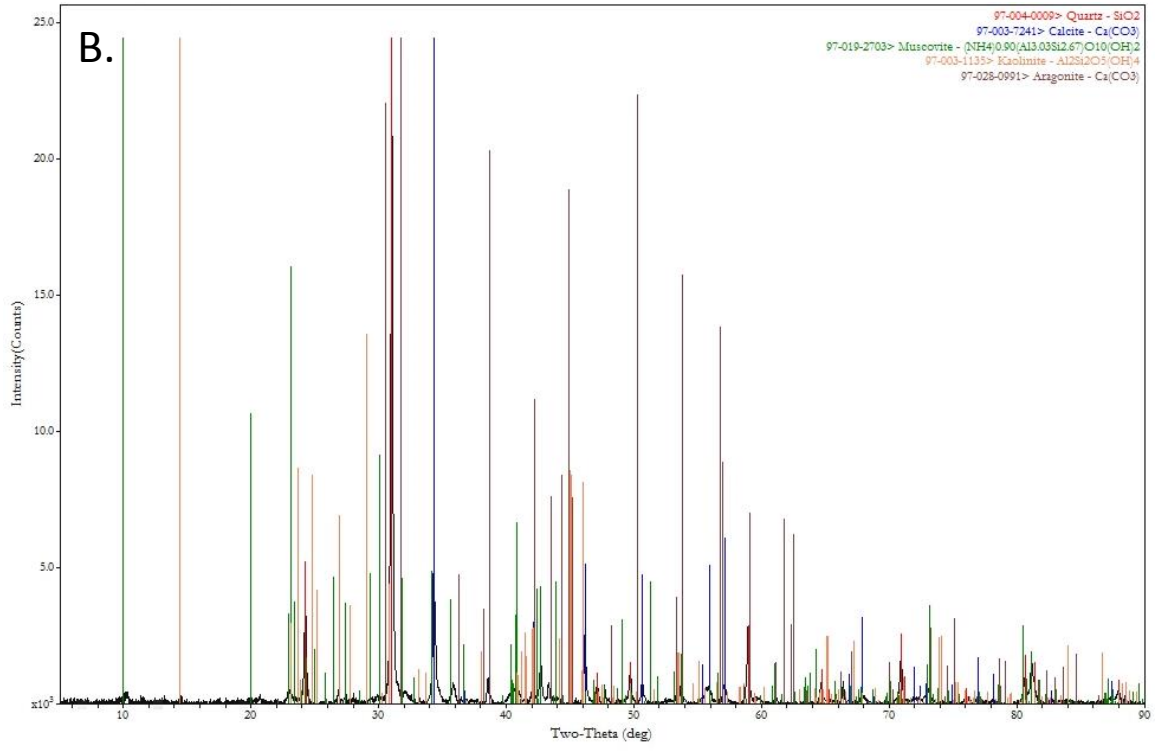


885

886 Figure SI.3 XRD spectra of the bulk mineralogy from drill cuttings from well 3's lateral bore A)
887 3550 m and B) 4250 m.

Table SI.1 The elemental composition of the FPW associated solids in mM as determined by alkaline fusion digestions and the associated relative standard deviation (RSD) as determined through three replicate digestions.

	S1	RSD	S2	RSD	S3	RSD	S3_Ox	RSD
Element	mM	%	mM	%	mM	%	mM	%
Mg	0.757	4.5	12.0	17.5	29.5	5.5	7.53	24.9
Al	10.7	30.0	96.9	3.8	57.0	8.8	2.17	37.2
Si	3,350	4.9	5,810	3.0	4,040	5.2	1,420	13.5
P	10.6	20.5	3.72	77.0	35.2	9.9	11.9	51.8
S	322	4.8	778	3.8	851	6.7	73.1	20.4
K	17.9	14.4	43.3	5.7	45.7	10.2	2.77	57.5
Ca	99.7	2.4	139	6.1	634	5.0	370	15.1
Mn	NM		1.74	3.9	NM		0.834	16.2
Fe	5,890	4.5	1,630	5.4	3,010	4.2	3,480	15.2
Zn	NM		2.44	46.6	NM		1.56	25.7
Sr	121	3.8	717	2.4	330	6.1	44.3	18.0
Ba	166	4.4	49.0	6.3	468	6.3	43.5	20.3

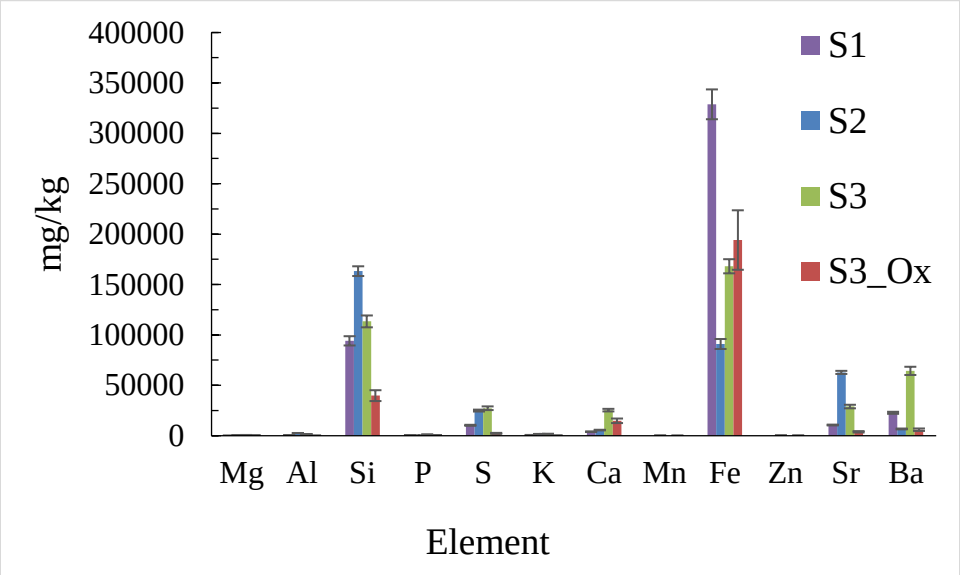


Figure SI.4 Elemental profile of FPW solids as determined through alkaline fusion digestion, in which the error bars represent ±1 standard deviation.

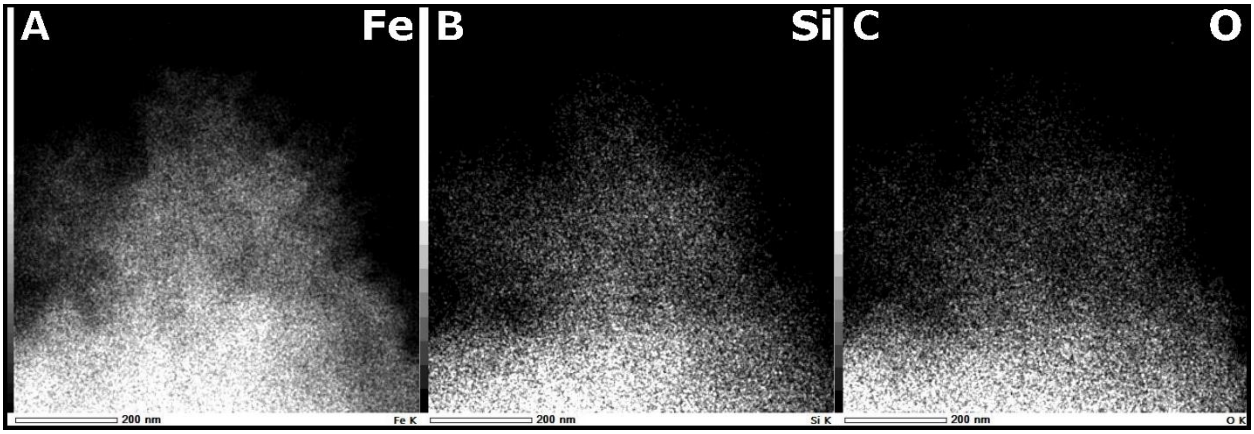


Figure SI.5 STEM elemental maps of particle agglomerates from S1 showing the distribution of A) Fe, B) Si, and C) O.

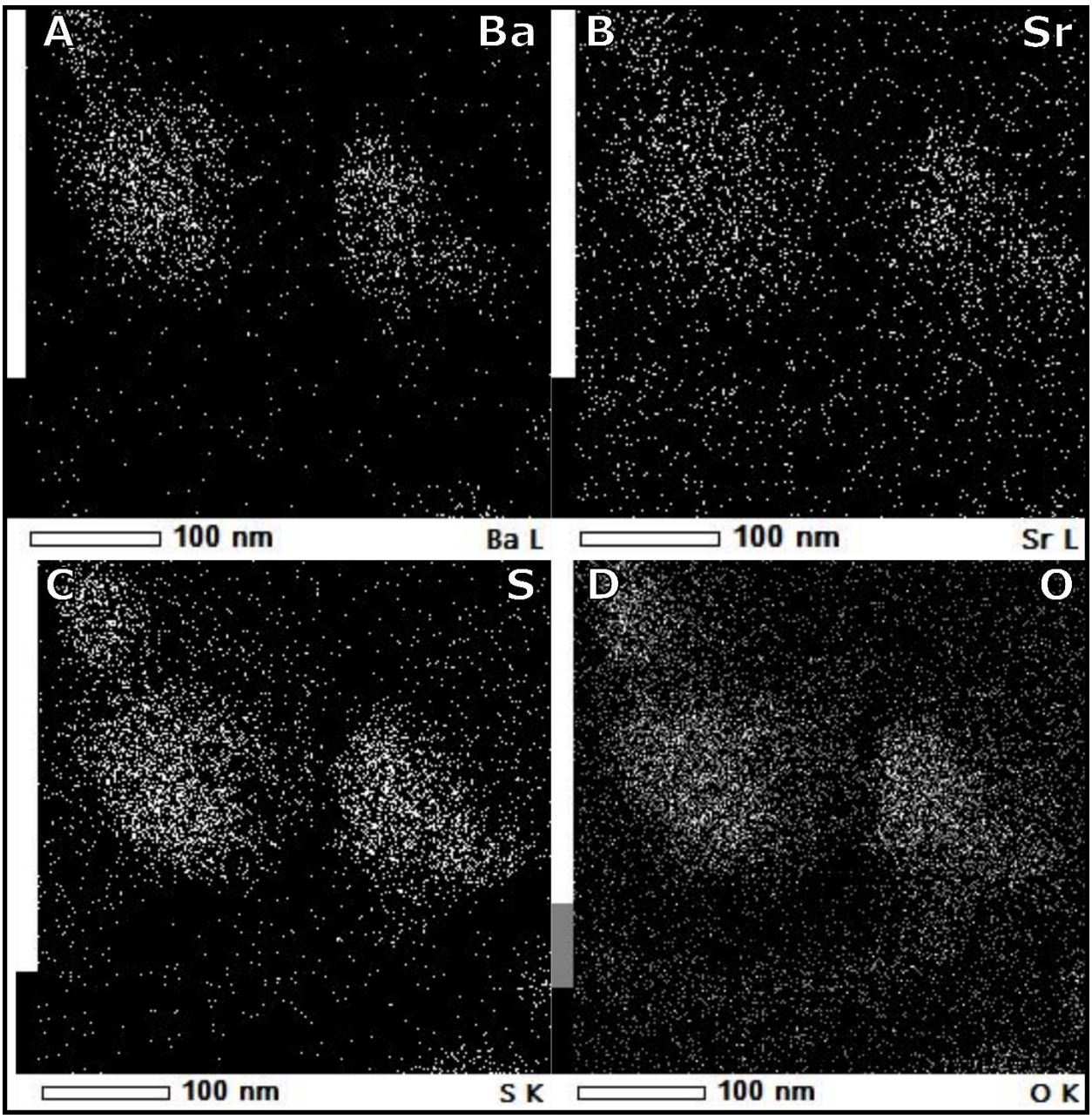
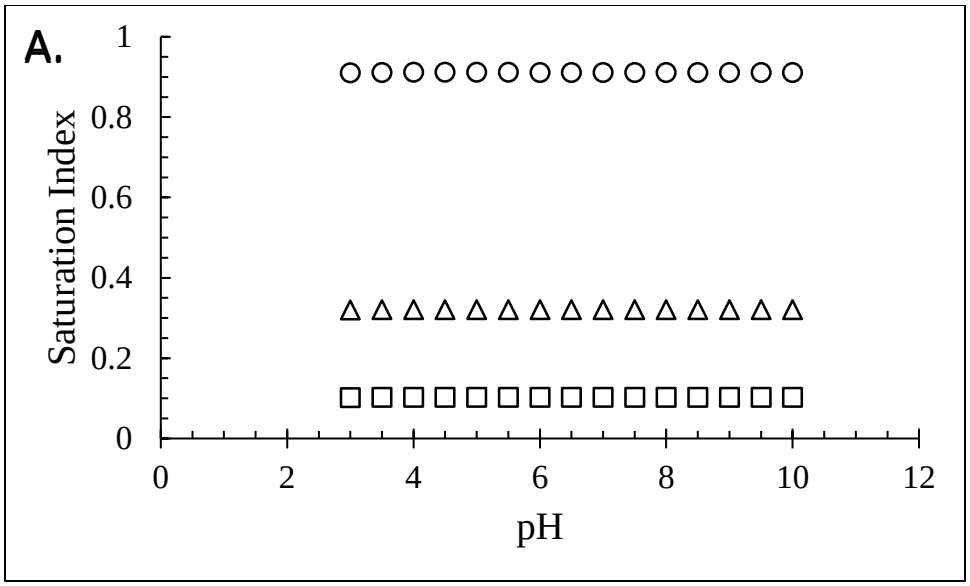
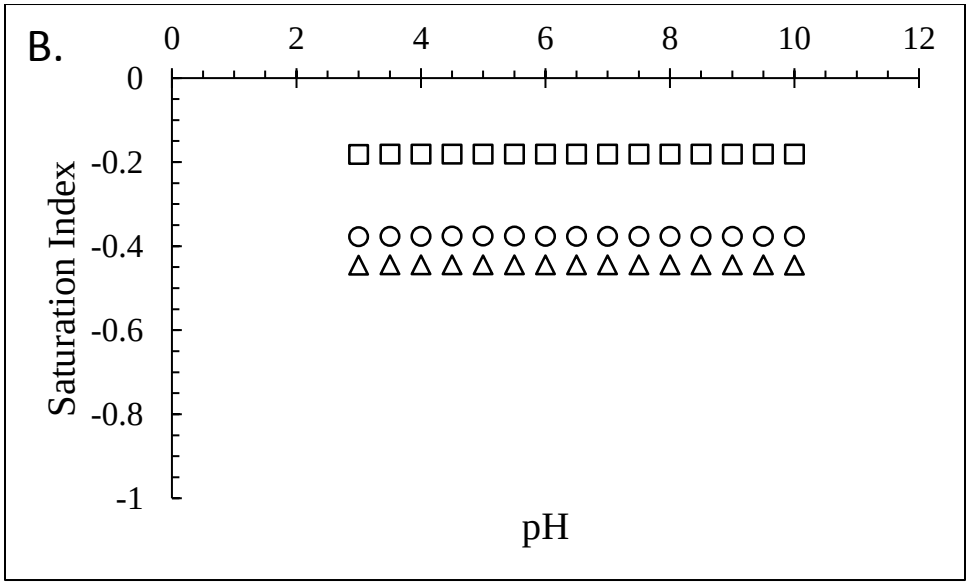


Figure SI.6 STEM elemental maps of particle agglomerates from S1 showing the distribution of A) Ba, B) Sr, C) S, and D) O.



904



905

906 Figure SI.7 Saturation indices of hypothesized minerals present FPW modeled using the element
907 data as a function of pH A) respect to barite, and B) celestine, with S1, S2, and S3 represented by
908 (□), (Δ), and (○), respectively.

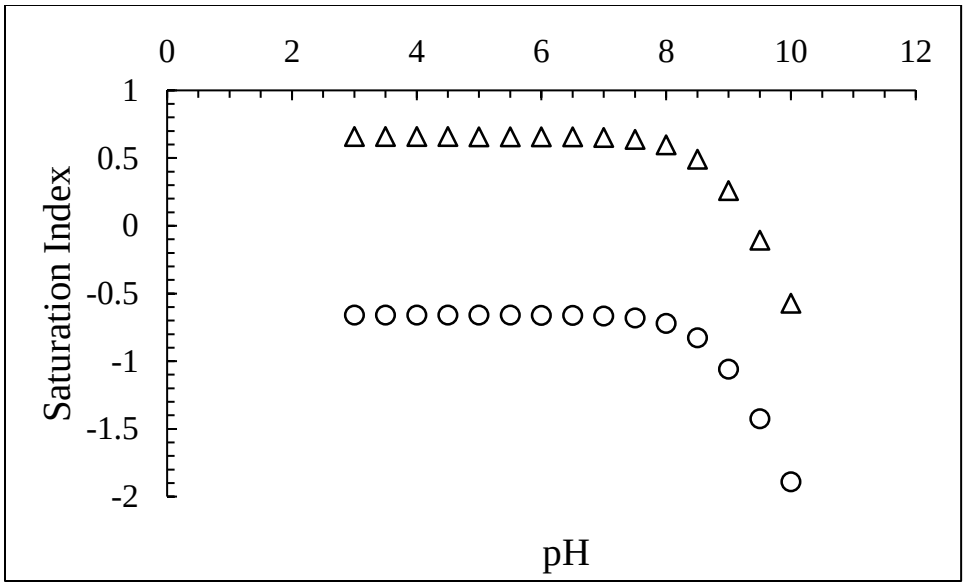


Figure SI.8 Saturation indices of potential silicate minerals present FPW modeled using the element data from S1 as a function pH for $SiO_{2(am)}$ (□), and quartz (Δ).

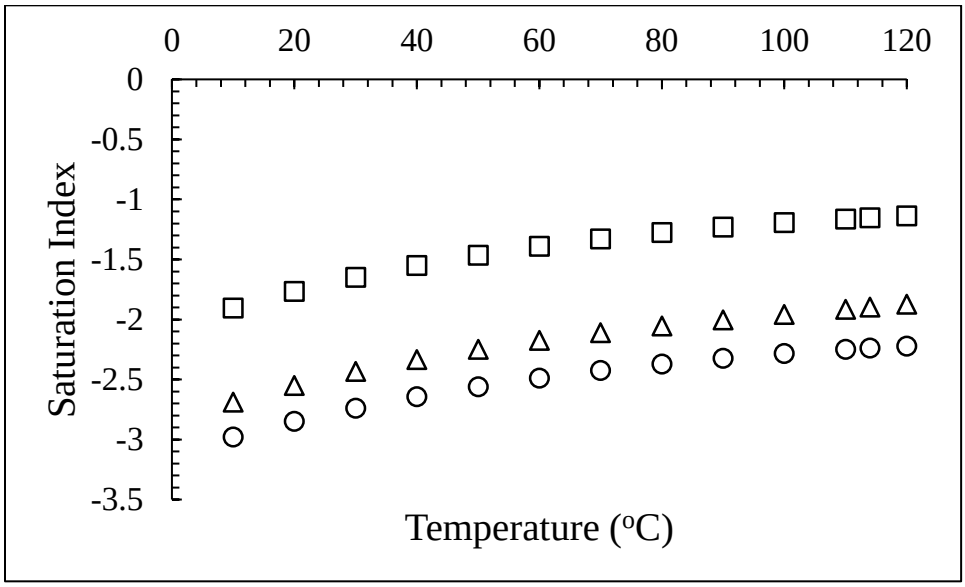
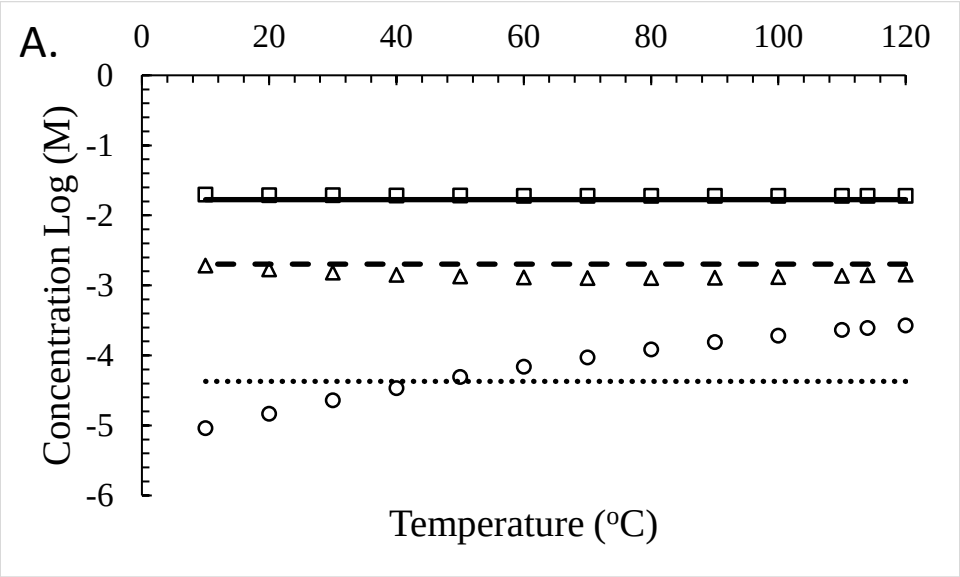
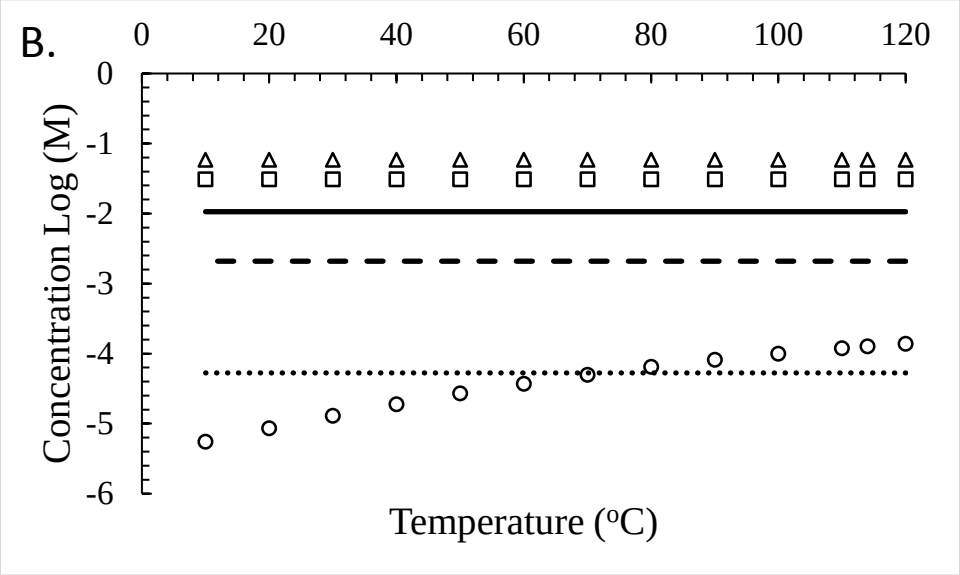


Figure SI.9 Saturation indices of CO_2 modeled using the FPW chemistry as a function of temperature with samples S1, S2, and S3 represented by (□), (Δ), and (○), respectively.

915


A.
0
20
40
60
80
100
120
0
-1
-2
-3
-4
-5
-6
Concentration Log (M)
Temperature (°C)

916


B.
0
20
40
60
80
100
120
0
-1
-2
-3
-4
-5
-6
Concentration Log (M)
Temperature (°C)

917

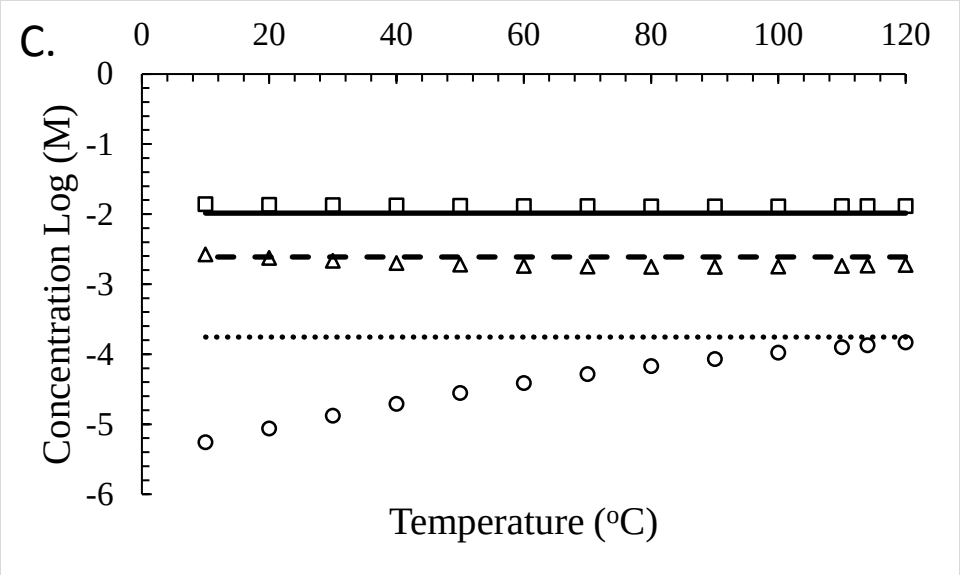


Figure SI.10 Comparison of the measured and model elemental concentrations in equilibrium with celestine and barite for A) S1, B) S2, and S3 in which the modeled Sr concentration is represented by (□), SO_4^{2-} by (Δ), and by Ba (○). The measured concentrations are for Sr, S, and Ba, represented by the solid, dashed and dotted lines respectively.

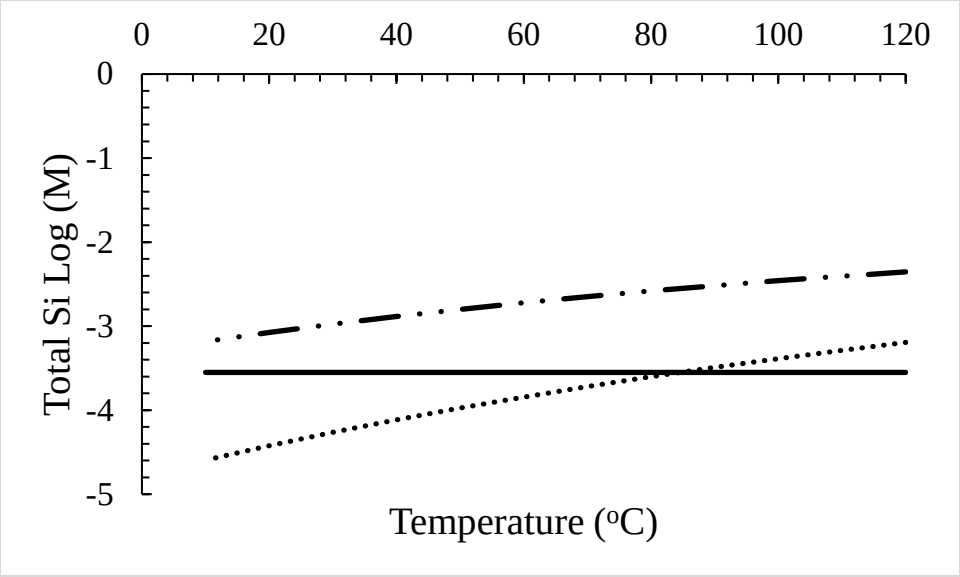


Figure SI.11 Comparison of the measured and model elemental concentrations of Si for S1 in equilibrium with silica minerals quartz and $SiO_{2(am)}$ as a function of cooling temperatures. The modeled concentrations of Si in equilibrium with quartz and $SiO_{2(a)}$ are represented by the dotted, and dotted and dashed lines, respectively. The solid line represents the measured Si concentration for S1.

Table SI.2 Data used to model the saturation indices for barite ($BaSO_4$), celestine ($SrSO_4$), quartz (SiO_2), and amorphous silica ($SiO_{2(am)}$) using FPW data from the Fayetteville, Bakken, Denver-Julesburg, Marcellus and Barnett Formations. NR denotes not reported and ND denotes not detected.

Fayetteville	Site	Ba (ppm)	Br (ppm)	Alkalinity as CO_3^{2-} (ppm)	Ca (ppm)	Cl (ppm)	Fe (ppm)	K (ppm)	Mg (ppm)	Mn (ppm)	Na (ppm)	SO_4^{2-} (ppm)	Si (ppm)	Sr (ppm)	pH
Warner et al.[5]	FS-1	5	96	1,136.6	221	5507	1	NR	56	2	3,232	ND	47	27	NR
	FS-2	4	122	486.0	345	10,165	13	NR	61	2	3,575	ND	13	14	NR
	FS-3	4	144	538.4	350	9,896	10	NR	75	3	4,607	ND	22	49	NR
	FS-4	3	101	874.6	386	10,312	1	NR	67	2	4,224	3	160	18	NR
	FS-5	3	97	800.0	284	6,771	8	NR	47	2	3,152	ND	18	26	NR
Bakken															
Shrestha et al.[6]	PW1	9.2	558	17.5	12,033	119,989	19.2	NR	1001	16.7	47,217	128	NR	774	NR
	PW2	12.4	384	84.5	8,573	75,892	30.2	NR	741	13.1	34,745	102	NR	551	NR
	PW3	26.3	91.6	428	372	21,728	0.7	NR	118	0.2	12,271	NR	NR	33.1	NR
	PW4	6.4	601	NR	15,346	136,220	22.3	NR	1299	15.8	60,571	293	NR	970	NR
	FW1	10.5	NR	145.35	9,683.3	118,666	96	NR	1273.3	7.1	61,466	650	NR	764	NR
Denver-Julesburg															
Rosenblum et al.[72]	HF	41.39	191.6	235.04	550	17,497	2.71	51.6	71.4	1.03	10,461	8.5	31.96	78	6.84

	VF1	41.07	244.4	41.97	1,081	24,955	6	75.9	119.3	0.19	14,215	>0.05	42	179	6.59
	VF2	31.55	265.1	67.15	1,204	27,103	4.8	65.4	130	0.33	14,794	>0.05	29	202	6.93
	VF3	14.15	81.3	164.88	365.5	12,724	19	52.9	40.5	0.29	7472	26.63	41	47	6.93
Marcellus															
Haluszczak et al.[8]		11,990	872	42.6	11,200	98,300	747	281	875	5.6	36,400	50	NR	2,330	6.2
Barbot et al.[9]		2,224	511	99	7,220	57,447	40.8	0	632	0	24,123	71	NR	1,695	6.56
He et al.[10]	A	730	NR	NR	2,170	29,000	NR	NR	249	NR	11,860	NR	NR	362	7.42
	B	236	NR	NR	15,021	104300	NR	NR	1720	NR	27946	14.8	NR	1799	6.4
Barnett															
Hayes and Severin[11]		3.6	589	441	1,600	34,700	24.9	316	255	0.86	18,850	709	NR	529	7.05

Warner et al.[5]: 5 Flowback water samples from the Fayetteville Formation.

Shrestha et al.[6]: 4 produced water and 1 flowback sample from the Bakken Formation.

Rosenblum et al.[7]: 4 total flowback samples, 1 from a horizontally fractured well and 3 from vertically fractured well from the Wattenberg field in the Denver-Julesburg Basin.

Haluszczak et al.[8]: median concentration from day 14 flowback from 7 wells in the Marcellus Formation.

Barbot et al.[9]: average from 95 Marcellus flowback samples.

He et al.[10]: two composite samples from separate wells in the Marcellus Formation. Sample A is a composite of days 1, 5 and 7, while sample B is a composite of FPW from days 1, 3, and 5.

941 Hayes and Severin[11]: the median of samples from 4 wells in the Barnett Formation from days 10-12.

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